

**COPY**

(12)

SECURITY CLASSIFICATION OF THIS PAGE

**UMENTATION PAGE**

Form Approved  
OMB No. 0704-0188

**AD-A233 089**

1b RESTRICTIVE MARKINGS

3 DISTRIBUTION/AVAILABILITY OF REPORT  
Approved for public release; distribution unlimited.

2b DECLASSIFICATION/DOWNGRADING SCHEDULE

Unclassified

4 PERFORMING ORGANIZATION REPORT NUMBER(S)

N00014-89-J-1237

5 MONITORING ORGANIZATION REPORT NUMBER(S)

6a NAME OF PERFORMING ORGANIZATION

Colorado State University

6b OFFICE SYMBOL  
(if applicable)

7a NAME OF MONITORING ORGANIZATION

6c ADDRESS (City, State, and ZIP Code)

Department of Chemistry  
Fort Collins, CO 80523

7b ADDRESS (City, State and ZIP Code)

8a NAME OF FUNDING/SPONSORING  
ORGANIZATION

Office of Naval Research

8b OFFICE SYMBOL  
(if applicable)

9 PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER

N00014-89-J-1237

8c ADDRESS (City, State, and ZIP Code)

800 North Quincy Street  
Arlington, VA 22217-5000

10. SOURCE OF FUNDING NUMBERS

PROGRAM  
ELEMENT NO

PROJECT  
NO

TASK  
NO

WORK UNIT  
ACCESSION NO

11 TITLE (Include Security Classification) Excited State Proton Transfer in the S<sub>1</sub> State of 2-Allylphenol, 2-Propenylphenol and 2-Propylphenol and their van der Waals Clusters with Water and Ammonia

12 PERSONAL AUTHOR(S)

S. K. Kim, S. C. Hsu, S. Li, E. R. Bernstein

13a TYPE OF REPORT

Technical Report

13b TIME COVERED

FROM \_\_\_\_\_ TO \_\_\_\_\_

14. DATE OF REPORT (Year, Month, Day)

March 7, 1991

15 PAGE COUNT

16 SUPPLEMENTARY NOTATION

17 COSATI CODES

FIELD

GROUP

SUB-GROUP

18 SUBJECT TERMS (Continue on reverse if necessary and identify by block number)

excited state intramolecular and intermolecular proton transfer, 2-allylphenol, 2-propenylphenol, 2-propylphenol, mass resolved excitation, threshold photo-ionization,

19 ABSTRACT (Continue on reverse if necessary and identify by block number)

SEE ATTACHED ABSTRACT

dispersed emission, intermolecular proton transfer, 2-allyl and 2-propenyl-phenol(NH<sub>3</sub>)<sub>n</sub>, dispersed emission spectra of 2-propylphenol (NH<sub>3</sub>)<sub>n</sub>

**DTIC**  
**SELECTED**  
**S**  
**D**  
**MAR 20 1991**

20 DISTRIBUTION/AVAILABILITY OF ABSTRACT

☒ UNCLASSIFIED/UNLIMITED ☐ SAME AS RPT ☐ DTIC USERS

21 ABSTRACT SECURITY CLASSIFICATION

Unclassified

22a NAME OF RESPONSIBLE INDIVIDUAL

Elliot R. Bernstein

22b TELEPHONE (Include Area Code)

(303) 491-6347

22c OFFICE SYMBOL

OFFICE OF NAVAL RESEARCH  
Contract N00014-89-J-1237  
TECHNICAL REPORT #71

Excited State Proton Transfer in the  
 $S_1$  State of 2-Allylphenol, 2-Propenylphenol  
and 2-Propylphenol and their van der Waals Clusters  
with Water and Ammonia

S. K. Kim, S. C. Hsu, S. Li, E. R. Bernstein

Submitted to the  
JOURNAL OF CHEMICAL PHYSICS

Chemistry Department  
Colorado State University  
Fort Collins, CO

March 7, 1991

Accession For	
NTIS	CRASH
DTIC	ENS
Unannounced	ED
Justification	
By	
Distribution /	
Availability Codes	
Dist	Avail. and/or Special
A-1	

Reproduction in whole or in part is permitted for  
any purpose of the United States Government.

This document has been approved for public release  
and sale; its distribution is unlimited.



## ABSTRACT

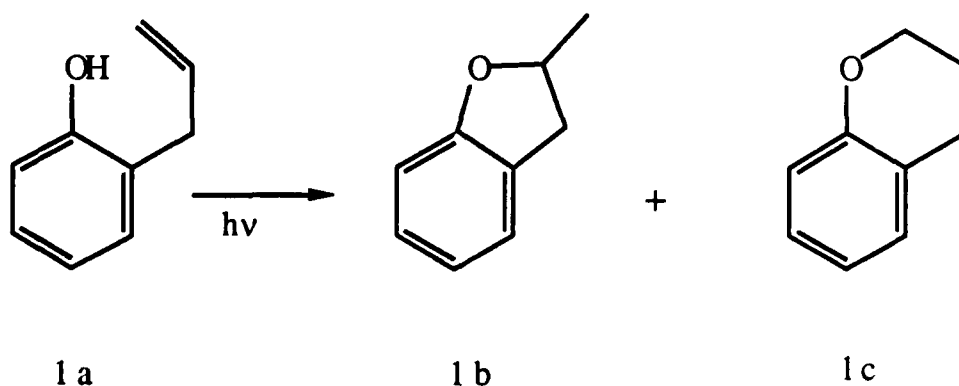
Excited state intramolecular and intermolecular proton transfer reactions in cold, isolated 2-allylphenol, 2-propenylphenol and 2-propylphenol and their clusters with water and ammonia are investigated employing a combination of spectroscopic techniques (mass resolved excitation, threshold photo-ionization, dispersed emission), a semi-empirical MNDO calculation (MOPAC 5) and a potential energy calculation of cluster structure. Threshold photo-ionization spectroscopy proves to be useful for the identification of molecular conformers in these systems but has mixed results for the identification of proton transfer in their clusters. The total collection of generated data suggests the following conclusions: 1. isolated, cold 2-allylphenol displays only one conformation which appears to have a significant stabilizing intramolecular interaction between the allyl group double bond and the hydroxyl group hydrogen atom; 2. 2-propenylphenol displays only one conformer; 3. 2-propylphenol has many conformations - probably more than five under the experimental conditions; 4. no evidence of intramolecular proton transfer can be found for these three isolated cold molecules; 5. no evidence for intermolecular proton transfer in water clusters has been found by any of the above techniques; and 6. evidence is found for intermolecular proton transfer in 2-allyl- and 2-propenyl-phenol( $\text{NH}_3$ ) $_n$ ,  $n \geq 3$ , in dispersed emission spectra. Dispersed emission spectra of 2-propylphenol( $\text{NH}_3$ ) $_n$ ,  $n \geq 3$  are too weak to yield conclusive evidence for intermolecular excited state proton transfer. Potential energy minimization calculations of cluster geometry suggest that the difference between water and ammonia cluster behavior with regard to proton transfer arises because water molecules hydrogen bond with the hydroxyl group (both  $\text{OH} \cdots \text{OH}_2$  and  $\text{HO} \cdots \text{HOH}$ ) and each other while ammonia molecules are more evenly distributed over the entire molecular structure of the phenol moiety. Apparently, for efficient proton transfer to occur in clusters the proton affinity of the solvent must be large and both the anion and the proton must be well solvated (stabilized) by the solvent.

## I. INTRODUCTION

The study of chemical reactions in cold, isolated van der Waals clusters has begun to make important contributions to the understanding of reaction mechanisms and dynamics<sup>1</sup>. Since cluster size and to some extent cluster structures can be controlled and/or selectively accessed in isolated clusters, the process of solvation and the effects of solvation on chemical reactions can be readily accessed in clusters.

A particularly interesting and apparently straightforward reaction for study in clusters is the elementary acid-base proton transfer reaction. Molecules which possess an aromatic hydroxyl group are attractive candidates for these reactions because they undergo large changes in  $pK_a$  upon photo-excitation from the ground ( $S_0$ ) to the first excited singlet ( $S_1$ ) state. For example, a Förster cycle<sup>2</sup> calculation for phenol gives  $pK_a(S_0) \sim 10.0$  and  $pK_a(S_1) \sim 3.6$ .<sup>3,4</sup> Substitution of halo, allyl and alkoxy groups on the phenol yield little change for these  $pK_a$  values while nitro-substituted phenols evidence increased acidity changes upon photo-excitation.<sup>3</sup>

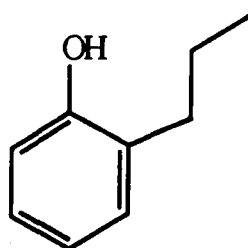
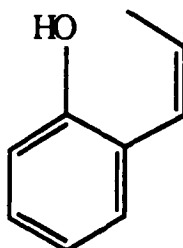
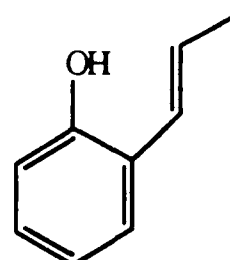
In ortho-substituted phenols both intra- and inter- molecular proton transfer can take place upon solution phase photo-excitation: 2-allylphenol (**1a**) provides a good example of such behavior. 2-allylphenol is known to photo-isomerize to generate cyclic ethers (**1b**, **c**) in non-polar solutions<sup>5-7</sup> as shown in Scheme I. The reaction is suggested to proceed via an intramolecular proton transfer between the hydroxyl hydrogen and the allyl group double bond.



Scheme I

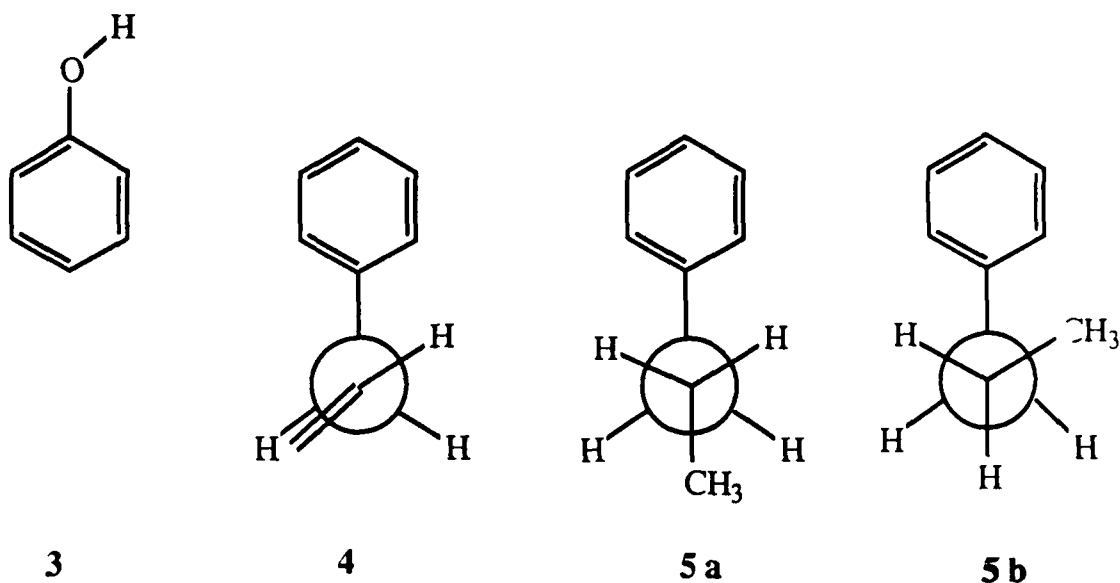
This photo-induced intramolecular proton transfer is suppressed in protic solvents:<sup>6</sup> intermolecular proton transfer prevails in protic solutions. If the solvent proton affinity could be carefully tuned or controlled, intra- and inter- molecular transfer might both be observed. While this possibility has not been realized in liquid solvation, the two different photo-induced proton transfer reactions might coexist under isolated conditions for appropriately structured clusters. Will 2-allylphenol undergo intra- or inter- molecular proton transfer in water and ammonia clusters? Can one determine if the transfer is intra- or inter- molecular?

The interaction between the phenol hydroxyl group and the 2-substituent on the ring, and thus the probability of intramolecular proton transfer, can be varied by changing the nature of the 2-substituent. Thus, 2-propenyl and 2-propylphenol can be investigated along with 2-allylphenol. In solution, no cyclization reaction is found for 2-propylphenol (**2a**) and 2-propenylphenol (**2b**, **2c**) undergoes cyclization (see Scheme I) only in the presence of a palladium catalyst<sup>8</sup>. The series of molecules **1a**, **2a**, **b**, **c** thus should allow one to distinguish between intra- and inter- molecular proton transfer in van der Waals clusters under controlled solvation conditions.

**2 a****2 b**  
syn**2 c**  
anti

The above 2-substituted phenols can be expected to exist in a number of different molecular conformations when cooled in a supersonic expansion: solute/solvent van der Waals clusters of a given mass containing these species will display multiple conformations, as well. Thus, intra- and inter- molecular proton transfer can be investigated as a function of molecular geometry, cluster geometry and number of solvent molecules present in the cluster.

The conformations of the related aromatic molecules containing only single substituents are already determined: 1. the OH group in phenol (3) is contained in the plan of the aromatic ring;<sup>9</sup> 2. the allyl group of allylbenzene (4) has the  $C_\alpha - C_\beta$  bond perpendicular to the plane of the aromatic ring and the allyl group double bond is eclipsed with one of the  $C_\alpha - H$  bonds;<sup>10</sup> 3. the entire heavy atom structure is planar for propenylbenzene ( $\beta$ -methylstyrene);<sup>11</sup> and 4. the propyl group in propylbenzene (5) has the  $C_\alpha - C_\beta$  bond perpendicular to the plane of the aromatic ring and the  $C_\beta - C_\gamma$  bond can be both anti and gauche to the ring.<sup>12</sup> These structures are expected to be altered to some degree in 2-substituted phenols due to steric and hydrogen bonding interactions.



Supersonic jet spectroscopy of molecules with multiple conformations becomes somewhat more complicated than usual. A number of techniques can be employed to distinguish different conformers of molecules from low lying vibronic features: 1. isotopic substitution will often yield large effects on vibronic features (e.g., shifts of greater than 5% in  $S_1$  vibrations) but only small changes in conformer origins;<sup>13</sup> 2. hole-burning experiments can be used to identify features ( $0_0^0$ ,  $X_0^1$ , etc.) common to a particular conformer;<sup>14</sup> and 3. ionization threshold measurements can demonstrate that conformers of a given molecule can have somewhat different ionization energies.<sup>15</sup>

The latter technique of threshold photo-ionization can also be important for cluster studies and the identification of chemical reactions in clusters.<sup>16-18</sup> Clusters of different structures can have different ionization energies (e.g., (pyrazine)<sub>2</sub>, (pyrimidine)<sub>2</sub>, etc.,<sup>19</sup> and naphthol/H<sub>2</sub>O and NH<sub>3</sub><sup>18</sup>). Clusters which have undergone proton transfer (e.g., naphthol/NH<sub>3</sub>) can also have a dramatically reduced ionization energy.<sup>16-18</sup>

This paper reports five separate studies on the 2-allyl-, 2-propenyl-, and 2-propyl- phenol molecules and their clusters with ammonia and water: mass resolved excitation spectroscopy; dispersed emission spectroscopy; threshold photo-ionization spectroscopy; semi-empirical MNDO calculation (MOPAC 5) of molecular structure; and cluster potential energy minimization calculations. Based on these results we determine that only intermolecular excited state ( $S_1$ ) proton transfer occurs for solute(NH<sub>3</sub>)<sub>n</sub>,  $n \geq 3$  clusters.

## II. EXPERIMENTAL PROCEDURES

A more detailed description of the supersonic jet apparatus and time of flight mass spectrometer employed in this effort can be found in earlier publications from our laboratory.<sup>20</sup> One-color mass resolved excitation spectroscopy (MRES) provides only crowded spectra for the samples used in this study probably due to extensive fragmentation from clusters of these molecules with impurity water. The ionization wavelength for two-color MRES is chosen to be low enough such that no cluster fragmentation is observed.

Threshold photo-ionization spectra are obtained with the  $S_1 \leftarrow S_0$  excitation energy ( $\nu_{ex}$ ) fixed for a particular transition and the  $I \leftarrow S_1$  ionization energy ( $\nu_{ion}$ ) scanned. The mass spectrometer extraction field for the created ions is  $\sim 100$  V/cm. The threshold photo-ionization value reported for the isolated molecules is obtained as follows: the spectrum is smoothed with Fourier transform filtering and the wavelength for the maximum signal derivative is found. The FWHM of the derivative function is reported as the width of the onset. Cluster signal intensities are smaller and the threshold spectra are quite broad: for these signals the detected onset of the ionization signal is reported for the threshold value.

2-allylphenol, 2-propenylphenol and 2-propylphenol are purchased from Aldrich Co. These samples are quite hygroscopic and contain ca. 2% water as an impurity. Two-color mass resolved excitation spectra of these species are not influenced by this difficulty: a problem, of course, arises for dispersed emission studies which are not mass selective. Samples are dried in vacuum over  $P_2O_5$ . Even for dried 2-propylphenol, the dispersed emission spectrum contains mostly 2-propylphenol/ $H_2O$  features. We are not able to obtain dispersed emission spectra of 2-propylphenol or 2-propylphenol/ammonia clusters free of water cluster contamination. The dispersed emission spectra for 2-allyl- and 2-propenyl-phenols and their ammonia clusters can be obtained free of interference from water cluster features.

The sample of 2-propenylphenol contains 80% anti- (**2c**) and 20% (**2b**) syn-isomers as determined by  $^{13}C$ - and  $^1H$ -NMR spectroscopy.

Samples are placed in the head of a pulsed nozzle and heated to ca. 40°C. 50 psi. He is typically employed as the expansion gas.

Fluorescence from these samples is weak so dispersed emission spectra are taken with UV cut-off filters placed in front of a photo-multiplier tube detector. The filters used are Hoya UV22, UV28, UV30, UV32, UV34, UV36, L38, and L40. This technique provides ca. 20 nm resolution dispersed emission spectra. The figure caption for the dispersed emission spectra of these systems contains transmission information for the filters.

Geometries for the 2-substituted phenols are calculated employing a semi-empirical MNDO algorithm, MOPAC 5 using PM3 and AM1 Hamiltonians.<sup>21</sup> Both calculations give the same results and thus only the PM3 results are reported herein. Starting geometries are varied to locate all minima on the potential surface for both the hydroxyl group and the hydrocarbon side chain. Stringent convergency criteria (GNORM = 0.01) still yield many minima for each molecule.

Cluster structures are also calculated employing the MOPAC 5 charges and minimum energy molecular geometries and a cluster energy minimization routine previously described.<sup>22</sup>



### III. RESULTS

#### A. 2-Allylphenol

##### 1. MRES and threshold photo-ionization of the isolated molecule

Two-color MRES of the 2-allylphenol  $S_1 \leftarrow S_0$  transition origin region is shown in Figure 1a. The lowest energy intense feature is at  $36029\text{ cm}^{-1}$  (peak A) and is assigned as a  $0_0^0$  transition for a conformer of 2-allylphenol. This origin has a  $50\text{ cm}^{-1}$  harmonic progression built on it. Figure 1b demonstrates that this  $50\text{ cm}^{-1}$  progression built on the 2-allylphenol origin must be due to motion of the allyl group because deuteration of the hydroxyl hydrogen does not change the progression spacing. Features labeled B and C in Figure 1 can be associated with low energy ring modes. No other electronic origins (conformers) have been identified for 2-allylphenol.

The  $I \leftarrow S_1$  photo-ionization spectrum for the peak A has a sharp ( $\Delta v_{\text{ion}} = 22\text{ cm}^{-1}$ ) onset at  $v_{\text{ion}} = 29662 \pm 4\text{ cm}^{-1}$  (Figure 2). These observations and assignments are summarized in Table I.

##### 2. MRES and threshold photo-ionization of water and ammonia clusters

Two-color MRES are displayed in Figure 3 for 2-allylphenol( $\text{H}_2\text{O}$ )<sub>1,2</sub>. The four different spectral regions displayed in Figure 3a can probably be associated with 2-allylphenol( $\text{H}_2\text{O}$ )<sub>1</sub> clusters of different structure. This point is best made by the threshold photo-ionization data presented in Table II. Most of the vibronic structure displayed in these spectra is due to allyl group motion: expansion of 2-allylphenol with  $\text{D}_2\text{O}$  does not change any of the observed spectra. The 2-allylphenol( $\text{H}_2\text{O}$ )<sub>2</sub> spectra displayed in Figure 3b are much simpler and show less structural diversity than do those of the 1:1 cluster. This observation will be rationalized later in light of cluster structure calculations and through comparison with results for 1-naphthol/water and ammonia clusters.<sup>18</sup>

The 2-allylphenol( $\text{H}_2\text{O}$ )<sub>1</sub> cluster absorbing ca.  $35,000\text{ cm}^{-1}$  (Figure 3a, top panel) undergoes almost 99% fragmentation to 2-allylphenol<sup>+</sup> and  $\text{H}_2\text{O}$  at the ionization energy employed ( $v_{\text{ion}} = 30,650\text{ cm}^{-1}$ ). The extensive fragmentation is observed until  $v_{\text{ion}}$  is lowered to  $28,608\text{ cm}^{-1}$  (lower than the bare molecule threshold).

Unlike the 2-allylphenol/water system for which intense cluster spectra are observed, 2-allylphenol/ammonia clusters yield only very weak one-color MRES and two-color MRES are too

weak to record. One-color MRES are very broad due to extensive fragmentation. We believe the reason ammonia cluster spectra are so weak is that most of the 2-allylphenol is clustered to impurity water. The water/allylphenol cluster binding energy is roughly twice that of ammonia/allylphenol.

### 3. Dispersed Emission Spectroscopy

Dispersed emission spectra of 2-allylphenol and its clusters with ammonia are presented in Figure 4. The dispersed emission from 2-allylphenol/water clusters is indistinguishable from that of the "bare molecule" (including, of course, a water contamination, since no mass resolution is employed). The dispersed emission from 2-allylphenol( $\text{NH}_3$ )<sub>n</sub> is shifted to lower energy: this red shift is usually indicative of an intermolecular excited state proton transfer. The shift in intensity in this instance is small because concentration of 2-allylphenol/ammonia clusters is small as shown above by one- and two- color MRES. The suggestion here is that the transfer is inter- and not intra- molecular because the red shifted emission is with respect to both bare molecule and water clusters. The more conclusive evidence with respect to these two possible proton transfer reactions will be provided through comparison between the behavior of the three different chromophores.

### 4. MOPAC 5 Calculations of Bare Molecule Structure

MOPAC 5 calculations for 2-allylphenol are presented in Table III and Figure 5. Fifteen locally stable conformations are calculated for this molecule. The three most stable conformations are shown in Figure 5. The single one observed probably corresponds to the lowest energy calculated structure. Interestingly, the syn-conformers ( $\tau_3 \sim 180^\circ$ ) tend in general to be more stable than the anti-conformers ( $\tau_3 \sim 0^\circ$ ). The more stable of the calculated conformers have an interaction between the hydroxyl hydrogen and the allyl group double bond.

### 5. Potential Energy Calculations of Cluster Structure

Calculation of 2-allylphenol/water and ammonia water cluster structures is not very reliable because the 2-allylphenol structure is not varied in the clustering process. Nonetheless, through comparison with results of phenol<sup>23</sup> and naphthol<sup>18</sup> clustered with water and ammonia, some general remarks can be made. First, most of the clustering takes place at the hydroxyl-allyl site.

Second, water molecules tend to bind to one another away from the allylphenol ring and side chain sites. Third, ammonia has a much more distributed geometry about the allylphenol because the ammonia-ammonia interaction (ca.  $500\text{ cm}^{-1}$ ) is smaller than the ammonia-allylphenol interaction (ca.  $900\text{ cm}^{-1}$ ). Fourth, detailed solute/solvent arrangements are dependent on the solute conformation chosen for clustering. Fifth, the solvent seems to position itself in many cluster configurations in such a fashion as to suggest that the original 2-allylphenol conformer structure would be altered to some extent by the solute/solvent interaction. And sixth, substantial hydrogen bonding takes place between the hydroxyl group and water in which the two possible hydrogen bonding configurations ( $\text{OH}\cdots\text{OH}_2$  and  $\text{HOH}\cdots\text{OH}$ ) are of roughly similar (ca.  $1600\text{ cm}^{-1}$ ) energy.

## B. 2-Propenylphenol

### 1. MRES and Threshold Photo-ionization of the Bare Molecule

The two-color MRES of 2-propenylphenol near the origin of the  $S_1 \leftarrow S_0$  transition is shown in Figure 6. The feature at  $32,725\text{ cm}^{-1}$  (A) is the lowest energy and most intense spectral feature and is assigned as the  $0_0^0$  transition for the anti-isomer (2c). Peaks marked B, C and E in Figure 6 are vibrations built on the A origin as can be seen from the threshold photo-ionization spectra presented in Figure 7. The sharp onset and large shift in ionization energy for peak D ( $33,491\text{ cm}^{-1}$ ) suggest that D is the  $0_0^0$  transition for the syn-isomer (2b). Table IV presents the threshold photo-ionization energies for these features: note that the values are the same for features A and D. The assignment of feature D is thus somewhat in question and data for the clusters can be employed to assist in the determination of the nature (i.e., vibronic feature associated with the A  $0_0^0$  transition, origin I, or  $0_0^0$  of the syn-isomer 2b).

### 2. MRES and Threshold Photo-ionization of Water and Ammonia Clusters

The two-color MRES of 2-propenylphenol/water and ammonia clusters are presented in Figure 8. The 1:1 cluster spectra are shifted from the isolated molecule spectra by  $-442$  and  $-750\text{ cm}^{-1}$ , respectively. These two clusters preserve the isolated molecule vibronic structure and Franck-Condon factors.

In the 2-propenylphenol( $\text{H}_2\text{O}$ )<sub>1</sub> and ( $\text{NH}_3$ )<sub>1</sub> cluster spectra, peak D (see Figure 6) is pronounced and unshifted (ca.  $0_0^0 + 768 \text{ cm}^{-1}$ ): this suggests that this feature (D in Figure 6) is vibronic in nature and not the syn-2-propenylphenol as given in **2b**.

2-propenylphenol( $\text{H}_2\text{O}$ )<sub>2</sub> and ( $\text{NH}_3$ )<sub>3</sub> spectra are broad and featureless as are spectra of higher order clusters in these series.

The threshold photo-ionization values for these clusters are given in Table II. The red shifts of the threshold photo-ionization values for water clusters appear to reflect only the binding energy differences between the cluster ground and excited states: no large, unexpected changes can be identified that would signify the onset of proton transfer. The behavior of 2-propenylphenol/ammonia clusters is quite similar. For comparison, the ionization energies for indole( $\text{H}_2\text{O}$ )<sub>1</sub>, indole( $\text{NH}_3$ )<sub>1,2</sub>, for which no proton transfer takes place, are shifted from the isolated indole molecule threshold photo-ionization energy by -3027, -4152, -4529  $\text{cm}^{-1}$ , respectively.<sup>15e</sup> One cannot determine if the  $\sim 600 \text{ cm}^{-1}$  red shift of the ionization threshold between 2-propenylphenol( $\text{NH}_3$ )<sub>2</sub> and ( $\text{NH}_3$ )<sub>3</sub> is due to proton transfer or simply the "usual" binding energy differences for the solvated ions and ground states. The ( $\text{NH}_3$ )<sub>3</sub> to ( $\text{NH}_3$ )<sub>4</sub> shift is comparable and cannot be employed to suggest proton transfer.

### 3. Dispersed Emission

The dispersed emission spectra of 2-propenylphenol and 2-propenylphenol( $\text{NH}_3$ )<sub>n</sub> are displayed in Figures 4c, d. The water cluster spectra are identical to the bare molecular emission spectrum. The large red shifted emission intensity for the ammonia clusters strongly suggests intermolecular proton transfer for both 2-allylphenol and 2-propenylphenol( $\text{NH}_3$ )<sub>n</sub>,  $n \geq 3$ , because (solution phase) intramolecular proton transfer does not take place in the latter system. No red shifted emission is found by exciting ( $\text{NH}_3$ )<sub>1</sub> (31,972  $\text{cm}^{-1}$ ) and ( $\text{NH}_3$ )<sub>2</sub> (31,846  $\text{cm}^{-1}$ ) clusters of 2-propenylphenol.

### 4. MOPAC 5 Calculation of Isolated Molecule Structure

MOPAC 5 PM3 calculation results for 2-propenylphenol are presented in Table V and Figure 9. For both the syn- and anti- conformers, the  $\tau_3 \sim 180^\circ$  geometry is more stable than the

$\tau_3 \sim 0^\circ$  geometry (see Table V for angle definitions). The anti,  $\tau_3 \sim 180^\circ$ ,  $\tau_1 \sim 0^\circ$  conformer (see Figure 9) is the most stable and probably the only one observed in the spectrum. Here too, interaction between the double bond and the hydroxyl group hydrogen is indicated.

## 5. Potential Energy Calculations of Cluster Structure

The results of these cluster calculations are similar to those reported allylphenol/water and ammonia. Hydrogen bonding is stronger for water clusters than ammonia clusters. Both  $\text{OH} \cdots \text{OH}_2$  and  $\text{HOH} \cdots \text{OH}$  bonding structures are obtained with the latter being ca.  $300 \text{ cm}^{-1}$  more tightly bound. The general findings parallel those described above for 2-allylphenol clusters.

### C. 2-Propylphenol

#### 1. MRES and Threshold Photo-ionization of the Bare Molecule

The two-color MRES of 2-propylphenol about its  $S_1 \leftarrow S_0$  origin region is presented in Figure 10 and its  $I \leftarrow S_1$  threshold photo-ionization spectrum is present in Figure 11 and summarized in Table VI. The features labeled I-V in these spectra and Table VI are most likely  $0_0^0$  transitions of five conformers of 2-propylphenol: origin transitions of various conformers appear in threshold photo-ionization as sharp features and vibronic states of conformers appear as broad features. Additionally, as given in Table VI, the features I-V have unique photo-ionization threshold energies. Firm assignment of the number of origins in this spectrum is not possible, but probably more than five (i.e., I-V and B) can be located.

#### 2. MRES and Threshold Photo-ionization of Water and Ammonia Clusters

The two-color MRES spectra of 2-propylphenol( $\text{H}_2\text{O}$ )<sub>1,2</sub> clusters are presented in Figures 12a and b. Only one-color MRES are obtained for ( $\text{NH}_3$ )<sub>n</sub> clusters due to their low concentration in the expansion. Due to the large number of possible 2-propylphenol molecule conformers, many cluster conformations are possible for each specific cluster mass. 2-Propylphenol( $\text{H}_2\text{O}$ )<sub>n</sub>,  $n = 1, 2$ , threshold photo-ionization data are presented in Table II. Ammonia cluster spectra are too weak to generate meaningful results for threshold photo-ionization.

### 3. Dispersed Emission

Almost all emission observed from 2-propylphenol samples, whether expanded with ammonia or not, comes from 2-propylphenol( $\text{H}_2\text{O}$ )<sub>n</sub>. Vacuum drying of the sample did not seem to change this significantly. Some changes in the red emission with high concentrations of ammonia may be present but they are difficult to quantify. The impurity water/2-propylphenol ratio is probably 1:1 and water binds to the phenol to exclude other solvents (binding energy for  $\text{H}_2\text{O} \sim 1600 \text{ cm}^{-1}$  and for  $\text{NH}_3 \sim 800 \text{ cm}^{-1}$ ).

### 4. MOPAC 5 Calculations of Molecular Structure

Seven different low energy structures are found for the 2-propylphenol system. The calculation still prefers syn- (OH) conformations over anti- (OH) conformations, but the energy differences are now quite small. The potential surface for this molecule is clearly very complex and the molecule is probably quite flexible with regard to hydroxyl and propyl group orientations and displacements.

No cluster detailed calculations are presented for this system. The general conclusions for cluster structure follow from the remarks made for the other two molecules in this study. Binding energies for water are ca.  $1500 \text{ cm}^{-1}$  and for ammonia ca.  $800 \text{ cm}^{-1}$ . Cluster formation can dramatically change and/or interconnect any of the local minima presented in Table VII because the potential energy surface around the local minima is so shallow.

## IV. DISCUSSION

### A. 2-Allylphenol

Experimental and calculational evidence presented in the last section is consistent with a reasonably strong internal hydrogen bond between the hydroxyl hydrogen and the terminal allyl group double bond (see Figure 5) yielding a low energy single 2-allylphenol molecular conformation. Note that 2-methyl-allylbenzene displays two conformers.<sup>24,25</sup>

The  $50 \text{ cm}^{-1}$  vibronic progression following the origin transition is not related to OH group motion, but to a change in orientation of the allyl group upon electronic ( $S_1 \leftarrow S_0$ ) excitation (see Figure 1). Most likely this change is associated with the large change in  $\text{pK}_a$  of the hydroxyl

proton upon electronic excitation and the concomitant increase in hydrogen bonding strength between the groups. Such a change is not observed in other allylbenzene species.<sup>10</sup>

The ionization potential for 2-allylphenol ( $65,690\text{ cm}^{-1}$ ) is not particularly low compared to those for 2-propenyl- and 2-propyl- phenol ( $63,155$  and  $66,100 \pm 100\text{ cm}^{-1}$ , respectively). Thus intramolecular proton transfer in bare molecule 2-allylphenol is not indicated by these data: a red shifted ionization threshold would be expected for internal proton transfer.

Dispersion emission for bare 2-allylphenol is also not particularly red shifted compared to that of the other species. We conclude that intramolecular proton transfer is absent in the bare 2-allylphenol molecule.

Neither water clusters nor small ( $n < 3$ ) ammonia clusters show dramatic red shifts for ionization thresholds or  $S_1 \rightarrow S_0$  emission: proton transfer (internal or external) does not occur for these systems either.

Red shifted emission is found for 2-allylphenol( $\text{NH}_3$ )<sub>n</sub>,  $n \geq 3$ . Since the red shift is similar to that reported for naphthol/ammonia systems<sup>1d,e</sup> and similar to that for 2-propenylphenol/ammonia clusters (no internal proton transfer identified for 2-propenylphenol even in solution), we conclude that intermolecular excited state proton transfer has occurred in 2-allylphenol ( $\text{NH}_3$ )<sub>n</sub>,  $n \geq 3$ .

As pointed out above and in a previous publication,<sup>18</sup> ammonia is much better at solvation of **both** the anion and the proton than is water. The interaction between ammonia and the chromophore is distributed over the entire cluster because ammonia is a (relatively) poor hydrogen bonder and the ammonia/ammonia interaction is only ca.  $500\text{ cm}^{-1}$ . Water, on the other hand, hydrogen bonds to the hydroxyl proton and itself (ca.  $1500\text{ cm}^{-1}$ ) and mostly avoids the remainder of the chromophore in its low energy cluster geometries.

#### B. 2-Propenylphenol

Calculated geometries for this molecule suggest that the propenyl double bond and the hydroxyl hydrogen interact only weakly. This is consistent with the absence of internal proton transfer in solution and absence of a vibrational progression in either propenyl or hydroxyl motion

following the  $0_0^0$  transition. All of this is consistent with threshold photo-ionization data: no special red shifts can be identified.

Cluster emission results are consistent with excited state proton transfer for  $n \geq 3$  in 2-propenylphenol/ammonia clusters. Again, only the dispersed emission evidences any proton transfer behavior. Comparisons with 2-allylphenol require the transfer in both instances to be intermolecular. Cluster structure calculations generate a reasonable explanation for the difference in behavior between water and ammonia clusters. Intermolecular proton transfer in the excited state is favored if both ions can be well solvated, as must happen in any event in solution due to solvent crowding and packing.

Threshold photo-ionization energies in this cluster system do not appear to change dramatically, over and above the general clustering trends, for  $n \geq 3$  ammonia clusters; however, emission data for these clusters indicate intermolecular proton transfer has occurred. In general, one must employ a member of detection techniques to identify chemical reactions in clusters.

#### C. 2-Propylphenol

Cold, isolated 2-propylphenol can have many conformations: calculations of molecular structure also yield a number of nearly equivalent (e.g.,  $500 \text{ cm}^{-1}$ ) local energy minima on the ground state potential surface. Interactions between the propyl moiety and the hydroxyl group seem to be favored even though no side chain  $\pi$  - system is present.

Experimental difficulties with water contamination make ammonia clusters of 2-propylphenol difficult to detect but weak emission data suggest that proton transfer can occur for the larger ( $n \geq 3$ ) ammonia clusters of 2-propylphenol.

## VI. CONCLUSIONS

Based on two-color mass resolved excitation, dispersed emission, and threshold photo-ionization spectroscopies and semi-empirical MNDO and cluster potential energy calculations, for 2-allyl, 2-propenyl-, and 2-propyl- phenol and their clusters with water and ammonia, we can conclude the following:



1. no **intramolecular** proton transfer occurs in any (isolated molecule or cluster) of these systems;
2. no **intermolecular** proton transfer occurs for clusters with  $(\text{H}_2\text{O})_n$  or  $(\text{NH}_3)_{1,2}$ ;
3. **intermolecular** excited state proton transfer occurs for larger clusters ( $n \geq 3$ ) of ammonia with 2-allyl- and 2-propenyl- phenol;
4. **intermolecular** excited state proton transfer may occur for 2-propylphenol  $(\text{NH}_3)_n$   $n \geq 3$ , but impurity water clustering makes this determination less certain than for the other chromophore species;
5. in all three of these species the isolated molecules appear to have a significant interaction between the hydrocarbon side chain and the hydroxyl group hydrogen;
6. if this latter "hydrogen bonding" interaction is strong (e.g., 2-allyl- and 2-propenyl-phenol), only one molecular conformer is present in the expansion, but if it is weak (e.g., 2-propylphenol), the molecular potential energy surface has many accessible nearly equivalent energy minima;
7. for excited state proton transfer to occur a combination of adequate ion solvation (both proton and anion) and good solvent proton affinity is required; and
8. cluster chemical reactions may be detected by dispersed emission, threshold photo-ionization, and/or two-color mass resolved excitation spectroscopies depending on the system under investigation.<sup>18</sup>

Future studies with these systems will include time resolved dispersed emission spectroscopy and time resolved two-color mass detected spectroscopy in order to determine the kinetics and mechanisms for these intermolecular proton transfer reactions.

## REFERENCES

1. a. N. F. Scherer, C. Sipes, R. B. Bernstein, A. H. Zewail, *J. Chem. Phys.* **92**, 3239 (1990), D. H. Semmes J. S. Baskin, A. H. Zewail, *J. Chem. Phys.* **92**, 3359 (1990) and A. H. Zewail, *Science* **242**, 1645 (1988). References to earlier studies can be found in these three references.  
b. J. A. Syage, *J. Chem. Phys.* **92**, 1804 (1990) and J. Steadman, J. A. Syage, *J. Chem. Phys.* **92**, 4630 (1990).  
c. C. Wittig, S. Sharpe, R. H. Beaudet, *Acc. Chem. Res.* **21**, 341 (1988).  
d. O. Cheshnovski and S. Leutwyler, *J. Chem. Phys.* **88**, 4127 (1988).  
e. T. Dorz, R. Knochenmuss and S. Leutwyler, *J. Chem. Phys.* **93**, 4520 (1990).
2. T. Förster, *Chem. Phys. Lett.* **17**, 309 (1972).
3. J. F. Ireland and P. A. H. Wyatt, *Adv. Phys. Chem.* **12**, 131 (1976)
4. W. Bartok, P. J. Lucchesi and N. S. Snider, *J. Am. Chem. Soc.* **84**, 1842 (1962).
5. G. Frater and H. Schmid, *Hel. Chim. Act.* **50**, 255 (1967).
6. S. Houry, S. Geresh and A. Shani, *Israel J. Chem.* **11**, 805 (1973).
7. S. Geresh, O. Levy, Y. Markovits and A. Shani, *Tetrahedron* **31**, 2803 (1975).
8. N. R. Davies and A. D. DiMichiel, *Aust. J. Chem.* **26**, 1529, (1973).
9. a. H. Abe, N. Mikami, M. Ito, *J. Phys. Chem.* **86**, 1768 (1982).  
b. H. Abe, N. Mikami, M. Ito, Y. Udagawa, *J. Phys. Chem.* **86**, 2567 (1982) and *Chem. Phys. Lett.* **93**, 217 (1982).  
c. A. Oikawa, H. Abe, N. Mikami, M. Ito, *J. Phys. Chem.* **87**, 5083 (1983); **88**, 5180 (1984).  
d. H. Mizuno, K. Okuyama, T. Ebata, M. Ito, *J. Phys. Chem.* **91**, 5589 (1987).
10. P. J. Breen, E. R. Bernstein, J. I. Seeman and H. V. Secor, *J. Phys. Chem.* **93**, 6731 (1989).
11. H. S. Im, E. R. Bernstein, J. I. Seeman, *J. Chem. Phys.*, to be published

12. J. I. Seeman, H. V. Secor, P. J. Breen, V. H. Grassian and E. R. Bernstein, *J. Am. Chem. Soc.* **111**, 3140 (1989)
13. a. R. Nowak, J. A. Menapace and E. R. Bernstein, *J. Chem. Phys.* **89**, 1309 (1988).  
b. J. A. Warren, E. R. Bernstein and J. I. Seeman, *J. Chem. Phys.* **88**, 871 (1988).  
c. V. H. Grassian, E. R. Bernstein, H. V. Secor and J. I. Seeman, *J. Phys. Chem.* **93**, 3470 (1989).  
d. H. S. Im, V. H. Grassian and E. R. Bernstein, *J. Phys. Chem.* **94**, 222 (1990).  
e. H. S. Im, E. R. Bernstein, J. I. Seeman, and H. V. Secor, *J. Am. Chem. Soc.*, accepted.  
f. J. I. Seeman, H. V. Secor, H. S. Im and E. R. Bernstein, *J. Am. Chem. Soc.* **112**, 7073 (1990).
14. R. J. Lipert and S. D. Colson, *J. Phys. Chem.* **93**, 3894 (1989).
15. a. M. A. Duncan, T. G. Dietz and R. E. Smalley, *J. Chem. Phys.* **75**, 2118 (1981).  
b. J. Hager, M. A. Smith and S. C. Wallace, *J. Chem. Phys.* **83**, 4820 (1985); **84**, 6771 (1986).  
c. J. Hager, M. A. Smith and S. C. Wallace, *J. Chem. Phys.* **89**, 3833 (1985).  
d. J. Hager, M. Ivanko, M. A. Smith and S. C. Wallace, *Chem. Phys. Lett.* **113**, 503 (1985).  
e. J. Hager, M. Ivanko, M. A. Smith and S. C. Wallace, *Chem. Phys.* **105**, 397 (1986).  
f. J. Hager, G. W. Leach, D. R. Demmer and S. C. Wallace, *J. Phys. Chem.* **91**, 3750, (1987).
16. D. Solgadi, C. Juvet and A. Tramer, *J. Phys. Chem.* **92**, 3313 (1988).
17. C. Juvet, C. Lardeux-Dedonder, M. Richard-Viard, D. Solgadi and A. Tramer, *J. Phys. Chem.* **94**, 5041 (1990).
18. S. K. Kim, S. Li and E. R. Bernstein, *J. Chem. Phys.* (submitted).

19. a. J. Wana and E. R. Bernstein, *J. Chem. Phys.* **84**, 927 (1986).  
b. J. Wana, J. A. Menapace and E. R. Bernstein, *J. Chem. Phys.* **85**, 777 (1986).  
c. J. Wana, J. A. Menapace and E. R. Bernstein, *J. Chem. Phys.* **85**, 1795 (1986).  
d. J. Wana and E. R. Bernstein, *J. Chem. Phys.* **85**, 3243 (1986).
20. E. R. Bernstein, K. Law and M. Schauer, *J. Chem. Phys.* **80**, 207 (1984).
21. J. J. P. Stewart, MOPAC, A General Molecular Orbital Package, 5th ed., (1988).
22. a. P. J. Breen, J. A. Warren, E. R. Bernstein and J. I. Seeman, *J. Chem. Phys.* **87**, 1972 (1987).  
b. J. A. Menapace and E. R. Bernstein, *J. Phys. Chem.* **91**, 2843 (1987).  
c. M. R. Nimlos, D. F. Kelley and E. R. Bernstein, *J. Phys. Chem.* **93**, 643 (1989).
23. S. Li, S. K. Kim, E. R. Bernstein, unpublished results.
24. T. Schaefer, R. Sebastian and T. A. Wildman, *Can. J. Chem.* **57**, 3005, (1979).
25. M. Oki and H. Iwamura, *Bull. Chem. Soc. Jpn.* **33**, 717 (1960).

**Table I**

Results of the excitation and the threshold ionization spectra for 2-allylphenol

Excitation Frequency ( $\nu_{\text{ex}}$ , $\text{cm}^{-1}$ )	$\text{I} \leftarrow \text{S}_1$ onset ( $\nu_{\text{ion}}$ , $\text{cm}^{-1}$ )	Onset width ( $\text{cm}^{-1}$ )	$\nu_{\text{ex}} + \nu_{\text{ion}}$ ( $\text{cm}^{-1}$ )	Excitation peak Assignment
36029 (peak A)	29662 $\pm$ 4	28	65691 $\pm$ 4	origin I
36079				origin I + $\nu_{\text{a}}$ (50 $\text{cm}^{-1}$ )
36129				origin I + 2 $\nu_{\text{a}}$
36179				origin I + 3 $\nu_{\text{a}}$
36142 (peak B)				origin I + $\nu_{\text{b}}$ (113 $\text{cm}^{-1}$ )
36192				origin I + $\nu_{\text{a}}$ + $\nu_{\text{b}}$
36242				origin I + 2 $\nu_{\text{a}}$ + $\nu_{\text{b}}$
36251 (peak C)				origin I + $\nu_{\text{c}}$ (222 $\text{cm}^{-1}$ )
36301				origin I + $\nu_{\text{a}}$ + $\nu_{\text{c}}$

Table II

Summary of Threshold Ionization of 2-allylphenol, 2-propenylphenol,  
2-propylphenol Clusters

Sample	Excitation Frequency ( $\nu_{\text{ex}}$ , $\text{cm}^{-1}$ )	$I \leftarrow S_1$ Threshold ( $\nu_{\text{ion}}$ , $\text{cm}^{-1}$ )	Ionization Potential (I.P.) ( $\nu_{\text{ex}} + \nu_{\text{ion}}$ , $\text{cm}^{-1}$ )	Shift of I.P. from baremolecule I.P. ( $\text{cm}^{-1}$ )
2-allylphenol	36029	29662 $\pm$ 4	65691 $\pm$ 4	0
2-allylphenol(H <sub>2</sub> O) <sub>1</sub>	35873	27268 $\pm$ 50	63141 $\pm$ 50	-2550
"	35853	26968 $\pm$ 50	62821 $\pm$ 50	-2870
2-allylphenol(H <sub>2</sub> O) <sub>2</sub>	36006	26968 $\pm$ 50	62974 $\pm$ 50	-2717
2-propenylphenol	32725	30430 $\pm$ 3	63155 $\pm$ 3	0
2-propenylphenol(H <sub>2</sub> O) <sub>1</sub>	32284	27380 $\pm$ 50	59664 $\pm$ 50	-3491
2-propenylphenol(NH <sub>3</sub> ) <sub>1</sub>	31972	26280 $\pm$ 50	58252 $\pm$ 50	-4903
2-propenylphenol(NH <sub>3</sub> ) <sub>2</sub>	31846	26380 $\pm$ 50	58226 $\pm$ 50	-4929
2-propenylphenol(NH <sub>3</sub> ) <sub>3</sub>	31861	25748 $\pm$ 50	57609 $\pm$ 50	-5546
2-propenylphenol(NH <sub>3</sub> ) <sub>4</sub>	31861	25228 $\pm$ 50	57089 $\pm$ 50	-6066
2-propylphenol	35953@	30200 $\pm$ 3	66153 $\pm$ 3	0
2-propylphenol(H <sub>2</sub> O) <sub>1</sub>	35607	26380 $\pm$ 50	61987 $\pm$ 50	-4166
"	35871	26210 $\pm$ 50	62081 $\pm$ 50	-4072
2-propylphenol(H <sub>2</sub> O) <sub>2</sub>	36092	27255 $\pm$ 50	63347 $\pm$ 50	-2806

@: The lowest energy origin

Table III

Conformational energies for 2-allylphenol calculated with MOPAC5/PM3

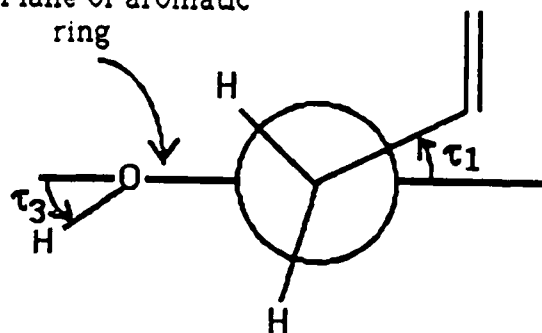
$\tau_3$ (degree)	$\tau_1$ (degree)	$\tau_2$ (degree)	$\Delta H_f$ (kcal/mol)
178	75	134	-10.4
169	-113	-127	-10.2
176	81	-133	-10.2
174	-38	-141	-9.88
165	-102	-97	-9.74
163	-113	120	-9.72
177	81	0	-9.30
179	65	-69	-9.19
167	-111	17	-8.91
3	95	-130	-8.62
1	84	134	-8.51
3	-13	-149	-8.44
3	26	131	-8.32
1	89	-7	-7.52
0	1	70	-7.37

Definition of  $\tau_1, \tau_2, \tau_3$  :

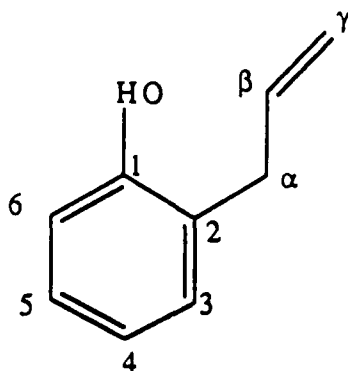
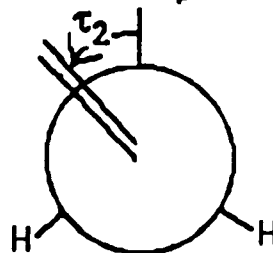
$\tau_1 = \text{angle}(\text{C}_2-\text{C}_3-\text{C}_\alpha-\text{C}_\beta)$ ,  $\tau_2 = \text{angle}(\text{C}_2-\text{C}_\alpha-\text{C}_\beta-\text{C}_\gamma)$ ,  $\tau_3 = \text{angle}(\text{C}_6-\text{C}_1-\text{O}-\text{H})$

Angles increase in the counterclockwise direction.

Plane of aromatic  
ring



Phenolic  
Group



**Table IV**

Results of the excitation and the threshold ionization spectra for 2-propenylphenol

Excitation Frequency ( $\nu_{\text{ex}}$ , $\text{cm}^{-1}$ )	$\text{I} \leftarrow \text{S}_1$ onset ( $\nu_{\text{ion}}$ , $\text{cm}^{-1}$ )	Onset width ( $\text{cm}^{-1}$ )	$\nu_{\text{ex}} + \nu_{\text{ion}}$ ( $\text{cm}^{-1}$ )	Excitation peak Assignment
32725 (peak A)	30430 $\pm$ 3	14	63155 $\pm$ 3	origin I
32912 (peak B)	30420 $\pm$ 3	16	63332 $\pm$ 3	origin I + $\nu_{\text{a}}$ (187 $\text{cm}^{-1}$ )
33008 (peak C)	30432 $\pm$ 10	120	63440 $\pm$ 10	origin I + $\nu_{\text{b}}$ (283 $\text{cm}^{-1}$ )
33491 (peak D)	29664 $\pm$ 3	13	63155 $\pm$ 3	origin I + $\nu_{\text{c}}$ (766 $\text{cm}^{-1}$ )
34004 (peak E)	30420 $\pm$ 20	250	64424 $\pm$ 20	origin I + $\nu_{\text{d}}$ (1279 $\text{cm}^{-1}$ )



**Table V**

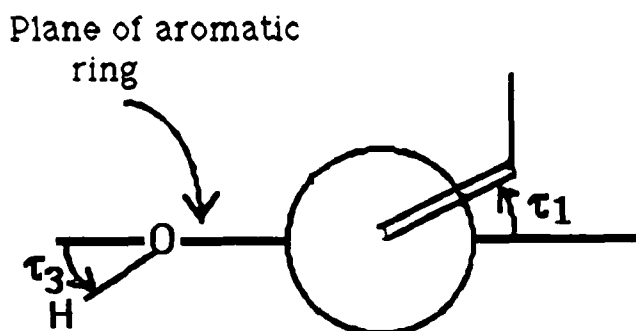
Conformational energies for 2-propenylphenol calculated with MOPAC5/PM3.

isomer	$\tau_3$ (degree)	$\tau_1$ (degree)	$\Delta H_f$ (kcal/mol)
syn	177	110	-14.0
	173	32	-13.8
	0	67	-12.5
	0	34	-12.2
	0	91	-12.1
	8	130	-9.7
anti	180	1	-16.3
	179	117	-15.2
	0	0	-14.5
	178	143	-14.3
	2	35	-14.1
	1	121	-13.0
	1	176	-12.5

Definition of  $\tau_1$ ,  $\tau_3$  :

$\tau_1$  = angle( $C_2-C_3-C_\alpha-C_\beta$ ),  $\tau_3$  = angle( $C_6-C_1-O-H$ ), see Table III for numbering.

Angles are increasing with counterclockwise direction.



**Table VI**

Results of the excitation and the threshold ionization spectra for 2-propylphenol

Excitation Frequency ( $\nu_{\text{ex}}$ , $\text{cm}^{-1}$ )	$\text{I} \leftarrow \text{S}_1$ onset ( $\nu_{\text{ion}}$ , $\text{cm}^{-1}$ )	Onset width ( $\text{cm}^{-1}$ )	$\nu_{\text{ex}} + \nu_{\text{ion}}$ ( $\text{cm}^{-1}$ )	Excitation peak Assignment
35953 (peak I)	30200 $\pm$ 3	17	66153 $\pm$ 2	origin I
35999 (peak A)	30208 $\pm$ 20	110	66207 $\pm$ 20	origin I + 46 $\text{cm}^{-1}$
36049 (peak II)	29953 $\pm$ 3	17	66002 $\pm$ 3	origin II
36189 (peak B)	29954 $\pm$ 2	17	66143 $\pm$ 3	? (a new origin or origin II + 140 $\text{cm}^{-1}$ )
36199 (peak III)	30001 $\pm$ 2	18	66200 $\pm$ 2	origin III
36226 (peak C)	29916 $\pm$ 3	14	66142 $\pm$ 3	? (a new origin or peak B + 47 $\text{cm}^{-1}$ )
36263 (peak D)	29950 $\pm$ 20	80	66213 $\pm$ 20	a vibration
36273 (peak IV)	29825 $\pm$ 2	13	66098 $\pm$ 2	origin IV
36288 (peak E)	29827 $\pm$ 20	110	66115 $\pm$ 20	origin IV + 15 $\text{cm}^{-1}$
36239 (peak V)	29722 $\pm$ 2	15	66051 $\pm$ 2	origin V

Table VII

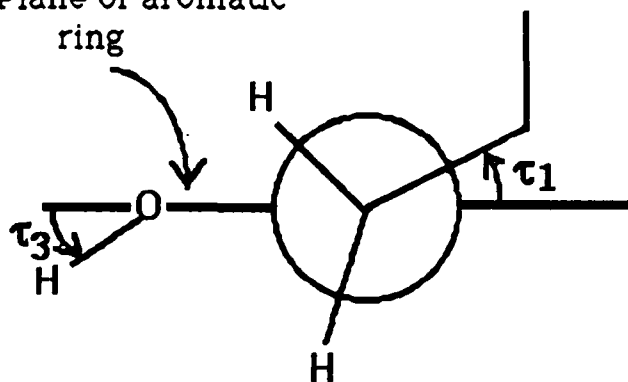
Conformational energies for 2-propylphenol calculated with MOPAC5/PM3.

=====

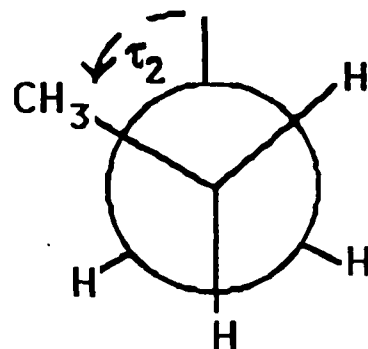
$\tau_3$ (degree)	$\tau_1$ (degree)	$\tau_2$ (degree)	$\Delta H_f$ (kcal/mol)
175	85	181	-41.0
181	76	72	-40.4
175	88	-75	-40.1
192	100	111	-39.8
2	83	181	-39.1
2	95	-75	-38.5
0	75	75	-38.2

=====

Plane of aromatic  
ring



Phenolic  
Group



### Figure Captions:

- Figure 1 2-color MRES of 2-allylphenol: a. Peak intensities are calibrated with dye curves of the excitation laser. Frequency of ionizing laser ( $\nu_{\text{ion}}$ ) is  $29800\text{ cm}^{-1}$ ; b. 2-color MRES of 2-allylphenol- $\text{d}_1$  for comparison with (a).
- Figure 2  $\text{I} \leftarrow \text{S}_1$  threshold photoionization spectrum of 2-allylphenol when the peak A is excited.
- Figure 3 2-color MRES of (a) 2-allylphenol( $\text{H}_2\text{O}$ ) $_1$  and (b) 2-allylphenol( $\text{H}_2\text{O}$ ) $_2$ . Peak intensities are calibrated with dye curves of the excitation laser. Energy of the ionizing laser ( $\nu_{\text{ion}}$ ) is  $30650\text{ cm}^{-1}$ .
- Figure 4 The dispersed fluorescence emission spectra for a) 2-allylphenol, b) 2-allylphenol( $\text{NH}_3$ ) $_n$ , c) 2-propenylphenol, d) 2-propenylphenol( $\text{NH}_3$ ) $_n$ . The excitation energies are also indicated, but the shapes of the spectra do not change upon changes in the excitation energy. The red-shifted emission is obvious in the ( $\text{NH}_3$ ) $_n$  clusters. The spectra of ( $\text{H}_2\text{O}$ ) $_n$  clusters are similar to the corresponding bare molecule spectrum in terms of wavelength distribution.

### Filter Transmission -

- 1) UV22-20% @ 200nm, 40% @ 220nm, 62% @ 240nm, 76% @ 260nm, 86% @ 280nm, 90% @ >300 nm.
- 2) UV28-0% @ <245nm, 7% @ 260nm, 44% @ 280nm, 77% @ 300nm, 87% @ 320nm, 90% @ >340nm.
- 3) UV30-0% <280nm, 38% @ 300nm, 83% @ 320nm, 90% >350nm.
- 4) UV32-0% <315nm, 2.5% @ 320nm, 52% @ 340nm, 80% @ 360nm, 84% @ 380nm, 87% @ >400 nm.
- 5) UV36-0% <340nm, 47% @ 360nm, 74% @ 380nm, 83% @ 400nm, 86% @ >420nm.
- 6) L38-0% <355nm, 47% @ 380nm, 78% @ 400nm, 86% @ 420nm, 90% @ >460nm.

7) L40-0% @ <370nm, 7% @ 380nm, 44% @ 400nm, 71% @ 420nm, 82% @ 400nm, 85% @ 460nm, 87% @ 480nm, 88% >500nm.

8) L42-0% <490nm, 44% @ 420nm, 83% @ 440nm, 87% >460nm.

Filters used for:

a. 2-allylphenol, 2-allylphenol(H<sub>2</sub>O)<sub>n</sub>, 2-allylphenol(NH<sub>3</sub>)<sub>n</sub>, 2-propylphenol(H<sub>2</sub>O)<sub>n</sub>– UV22, UV28, UV30, UV32, UV34, UV36, L38.

b. 2-propenylphenol, 2-propenylphenol(H<sub>2</sub>O)<sub>n</sub>, 2-propenylphenol(NH<sub>3</sub>)<sub>n</sub>– UV30, UV32, UV36, L38, L40, L42.

Figure 5 The three most stable conformations of 2-allylphenol calculated with MOPAC 5/PM3.

Figure 6 2-color MRES of 2-propenylphenol. Peak intensities are calibrated with dye curves of the excitation laser. Energy of ionizing laser ( $\nu_{\text{ion}}$ ) is 31380 cm<sup>-1</sup>.

Figure 7 I←S<sub>1</sub> threshold photoionization spectra of 2-propenylphenol. Excitation energy is written beside each spectrum.

Figure 8a,b: 2-color MRES of 2-propenylphenol(H<sub>2</sub>O)<sub>1</sub> and (H<sub>2</sub>O)<sub>2</sub>. Peak intensities are calibrated with dye curves of the excitation laser. Energy of ionizing laser ( $\nu_{\text{ion}}$ ) is 29000 cm<sup>-1</sup>.

c,d,e: 2-color MRES of 2-propenylphenol(NH<sub>3</sub>)<sub>1,2,3</sub>. Peak intensities are calibrated with dye curves of the excitation laser. Energy of the ionizing laser ( $\nu_{\text{ion}}$ ) is 27700 cm<sup>-1</sup>. The spectra for higher clusters are featureless and broad.

Figure 9 The most stable geometry of anti-2-propenylphenol (a) and the two most stable conformations of syn-2-propenylphenol (b,c).

Figure 10 2-color MRES of 2-propylphenol. Peak intensities are calibrated with dye curves of the excitation laser. Energy of ionizing laser ( $\nu_{\text{ion}}$ ) is 31600 cm<sup>-1</sup>.

Figure 11 I←S<sub>1</sub> threshold photoionization spectra of 2-propylphenol. Excitation energy is written beside each spectrum.

Figure 12a,b 2-color MRES of 2-propylphenol( $\text{H}_2\text{O}$ )<sub>1</sub> and ( $\text{H}_2\text{O}$ )<sub>2</sub>. Peak intensities are calibrated with dye curves of the excitation laser. Energy of ionizing laser ( $\nu_{\text{ion}}$ ) is 29960  $\text{cm}^{-1}$ .

c 1-color MRES of 2-propylphenol( $\text{NH}_3$ )<sub>1</sub>. Peak intensities are calibrated with dye curves of the excitation laser. The spectra for higher clusters are featureless and broad.

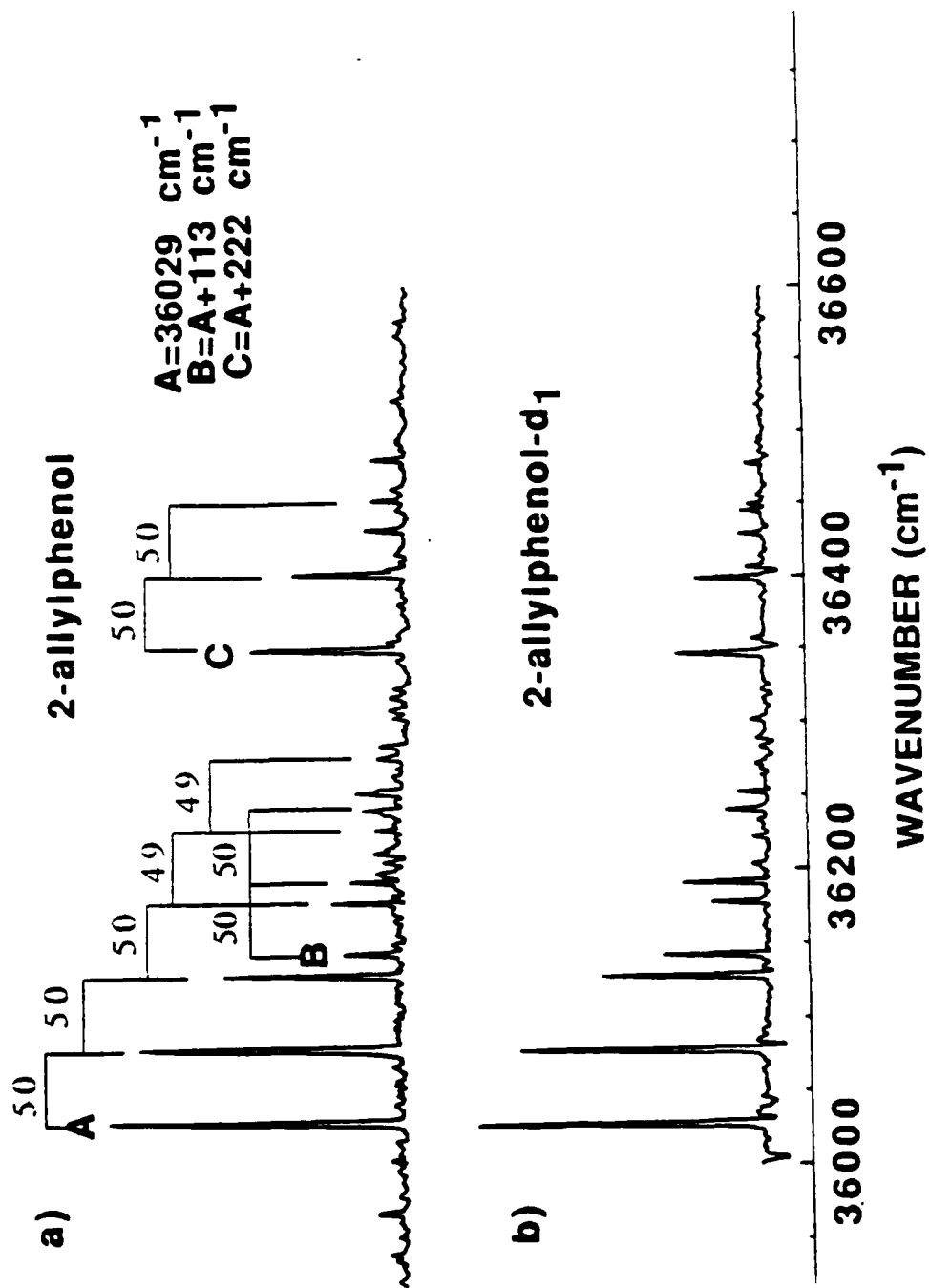


Figure 1

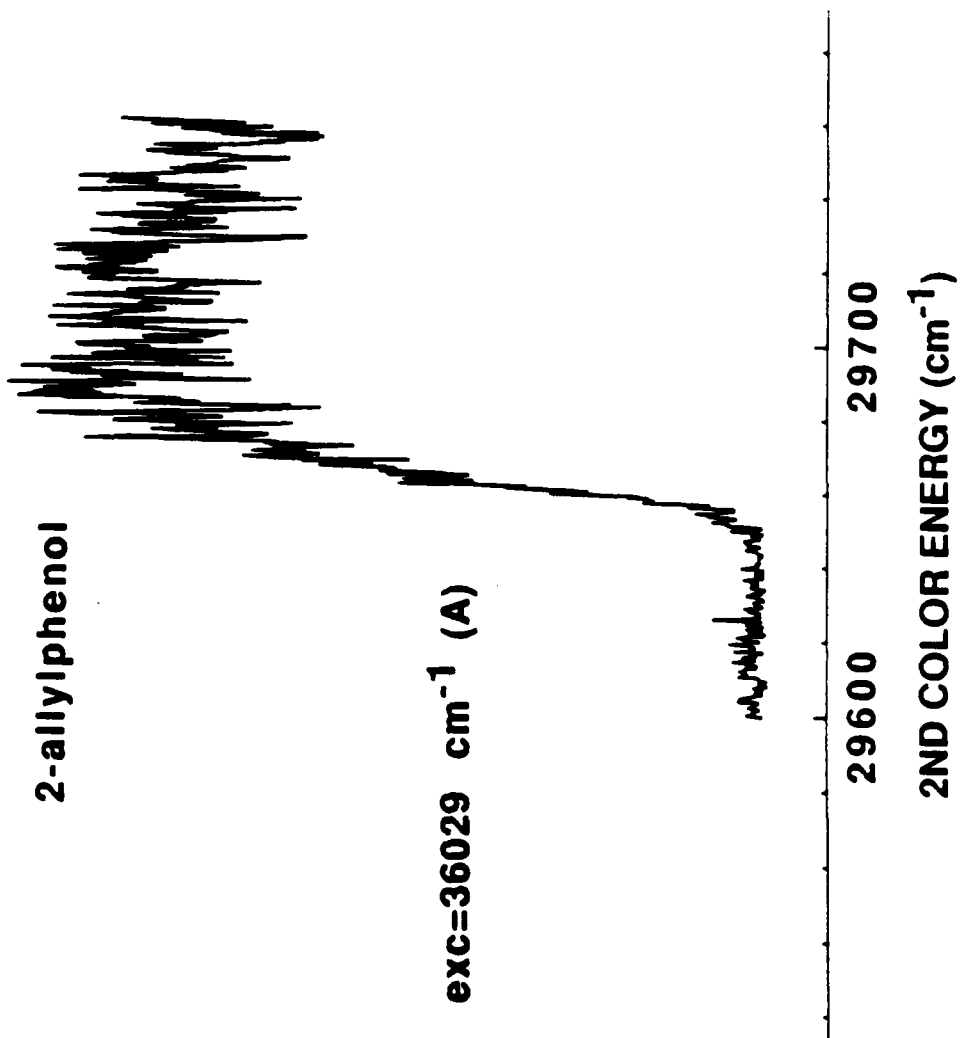


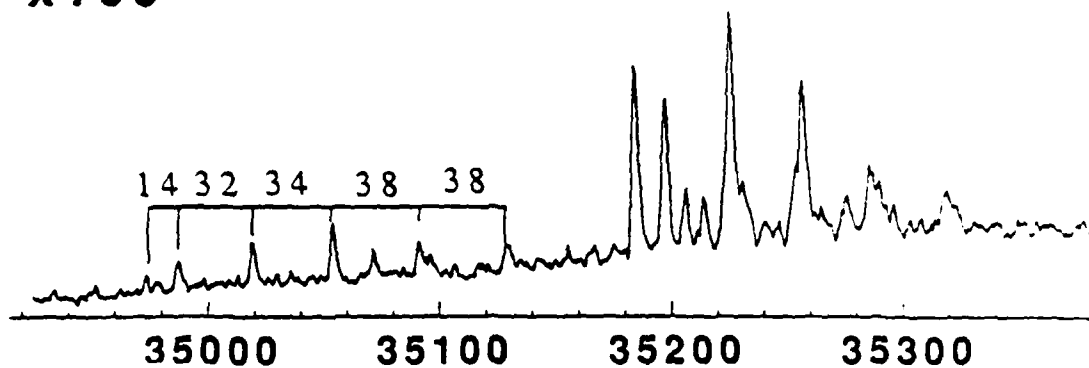
Figure 2



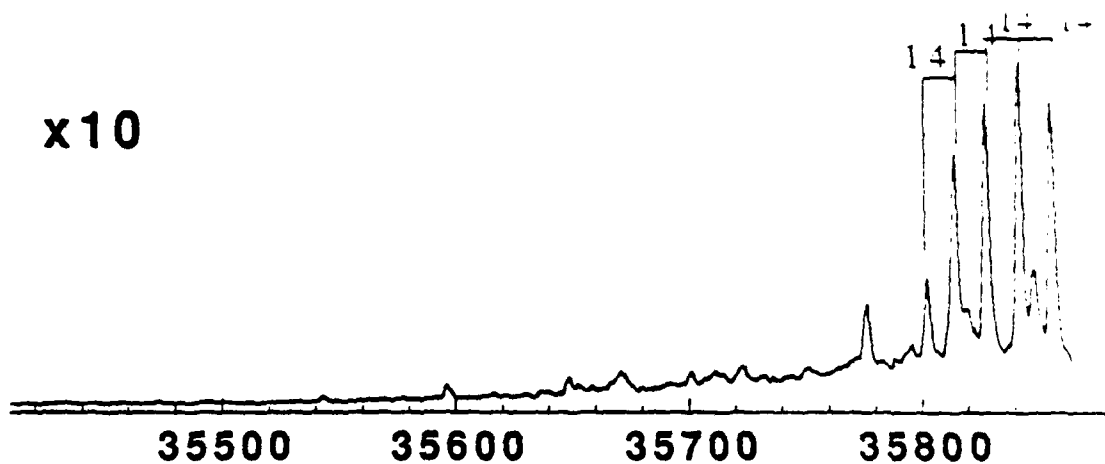
2-allylphenol( $\text{H}_2\text{O}$ )<sub>1</sub>

a)

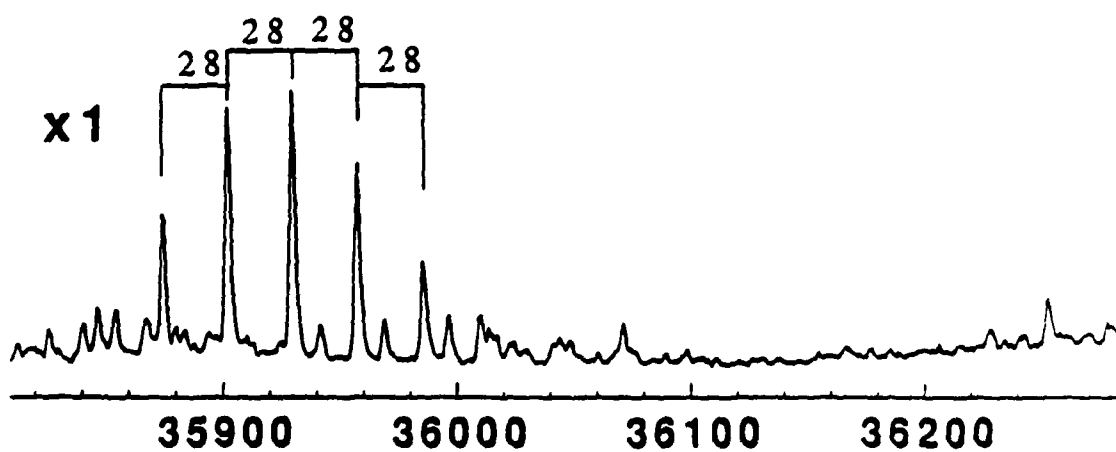
x100



x10



x1



WAVENUMBER ( $\text{cm}^{-1}$ )

Figure 3a

b)

2-allylphenol( $\text{H}_2\text{O}$ )<sub>2</sub>

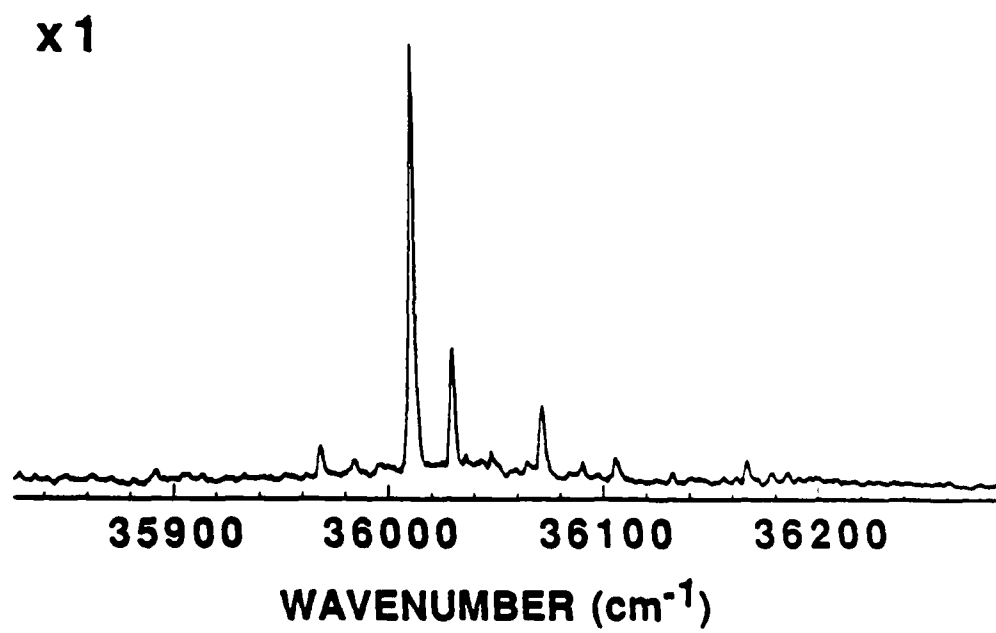
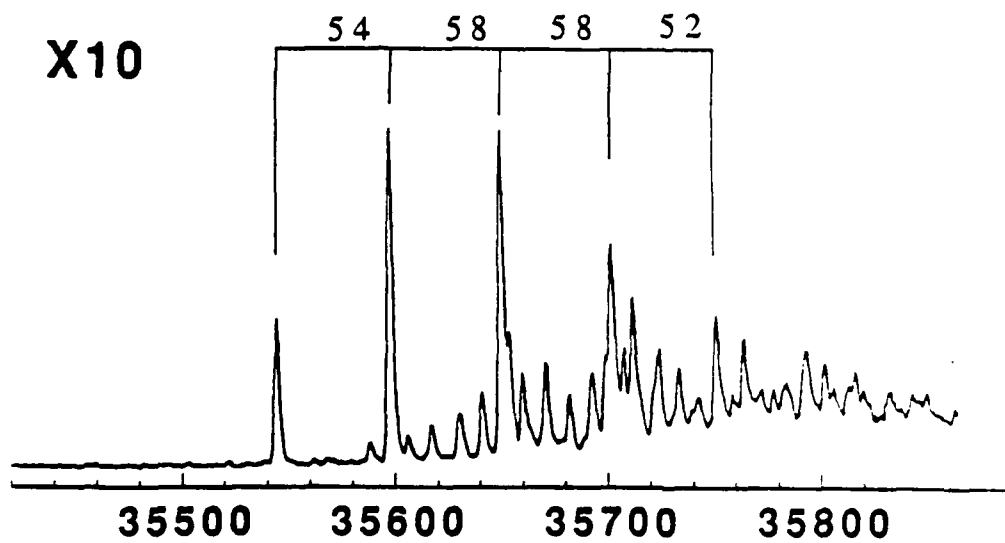


Figure 3b

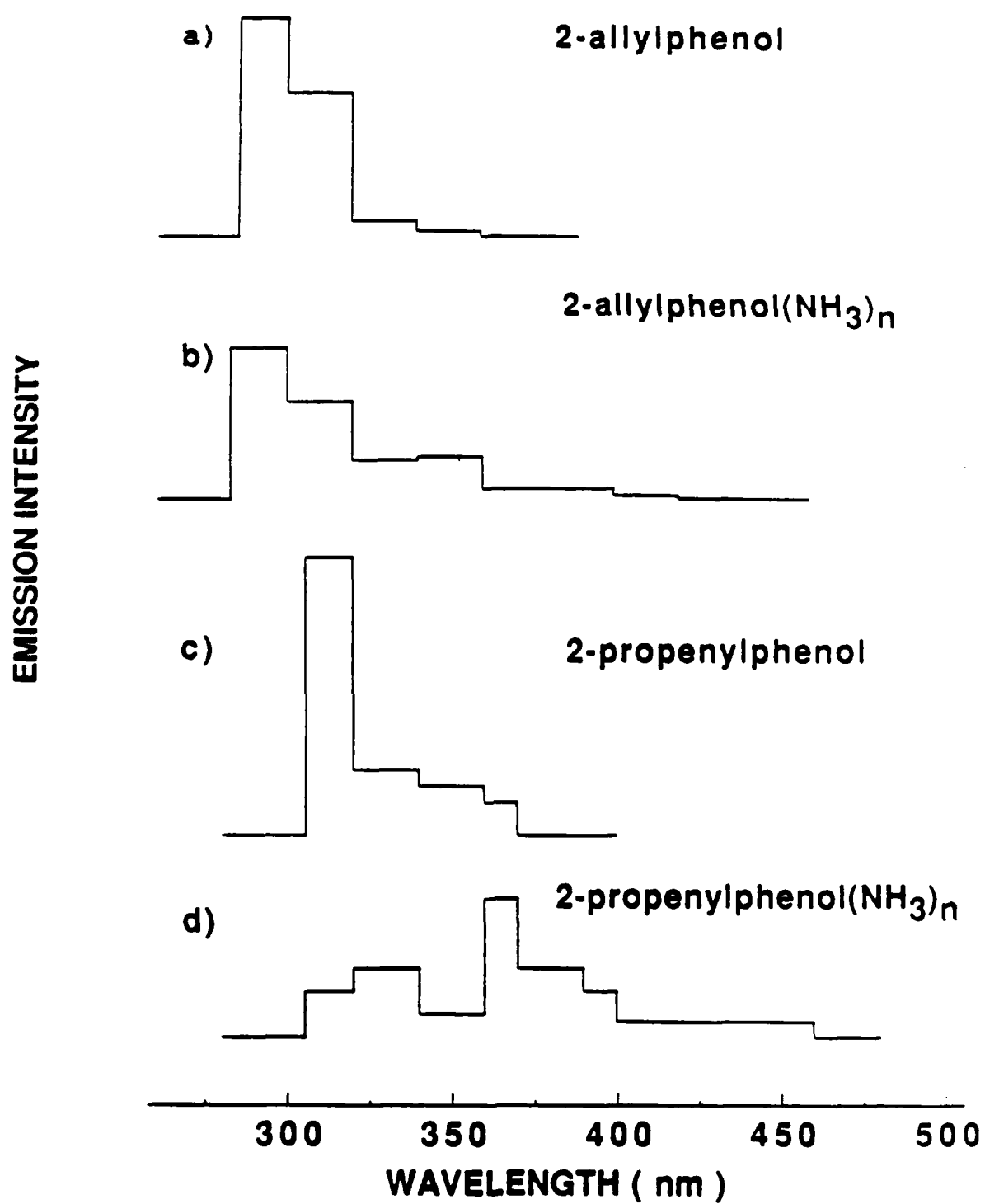


Figure 4

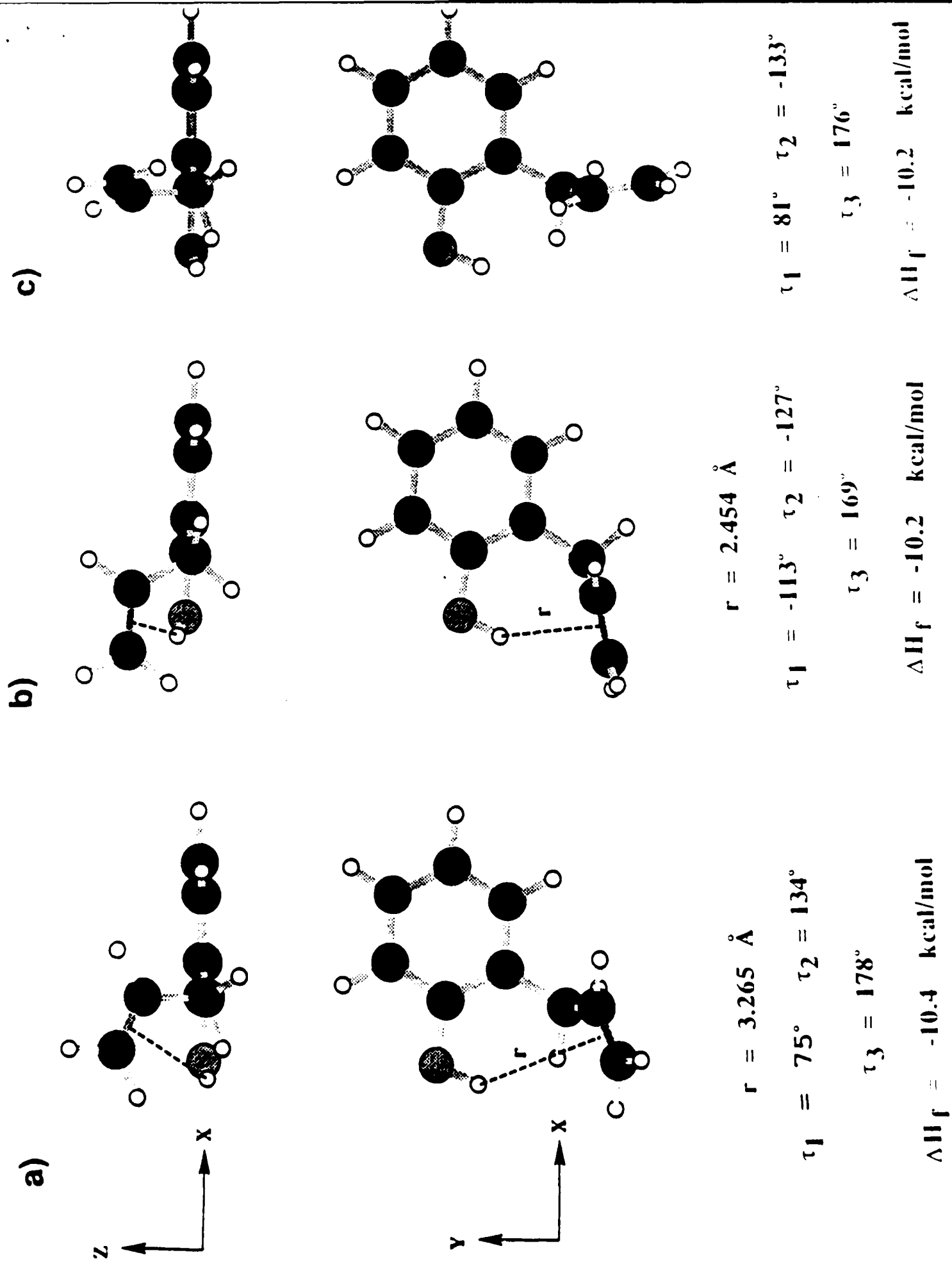


Figure 5

2-propenylphenol

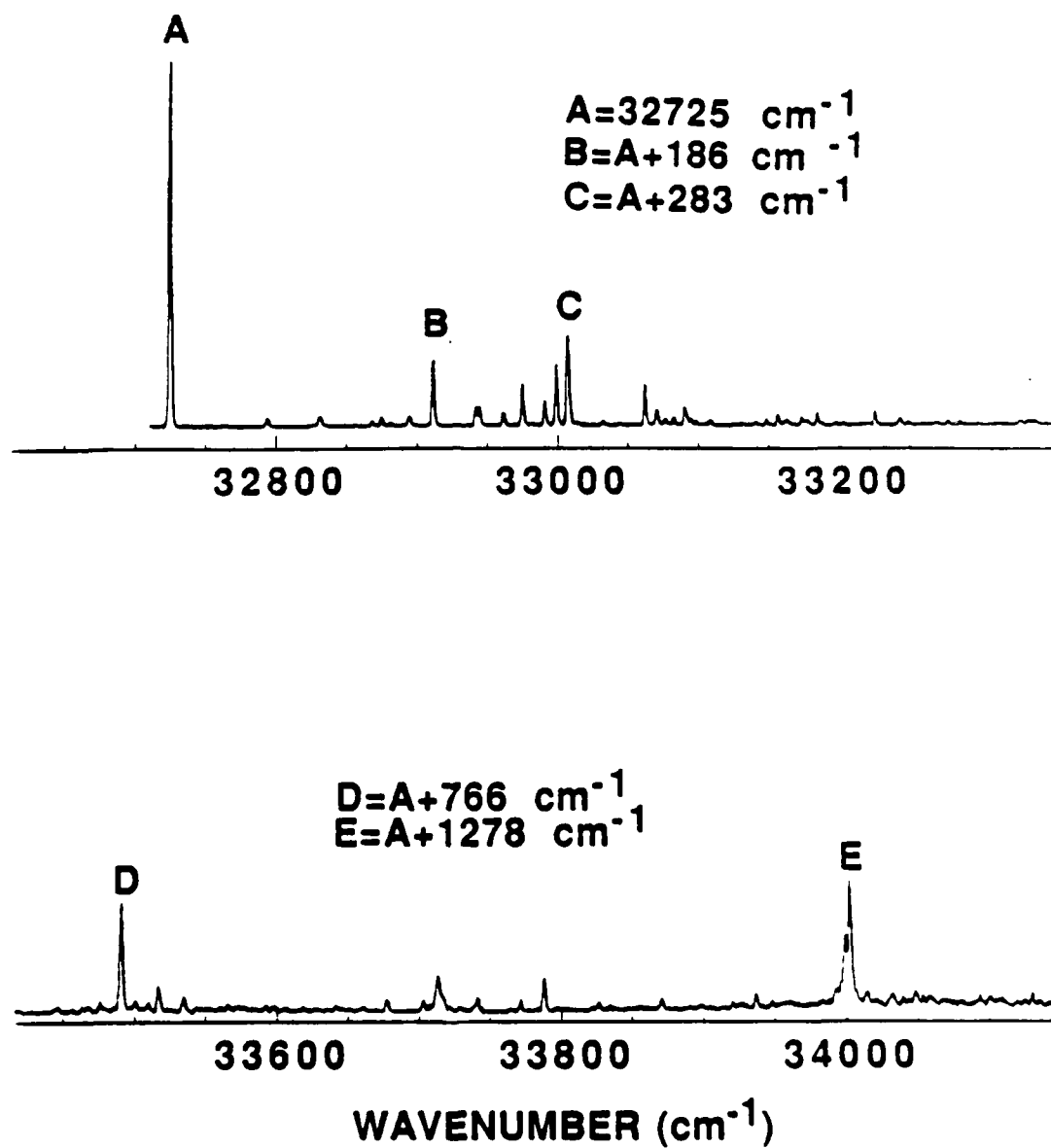


Figure 6

2-propenylphenol

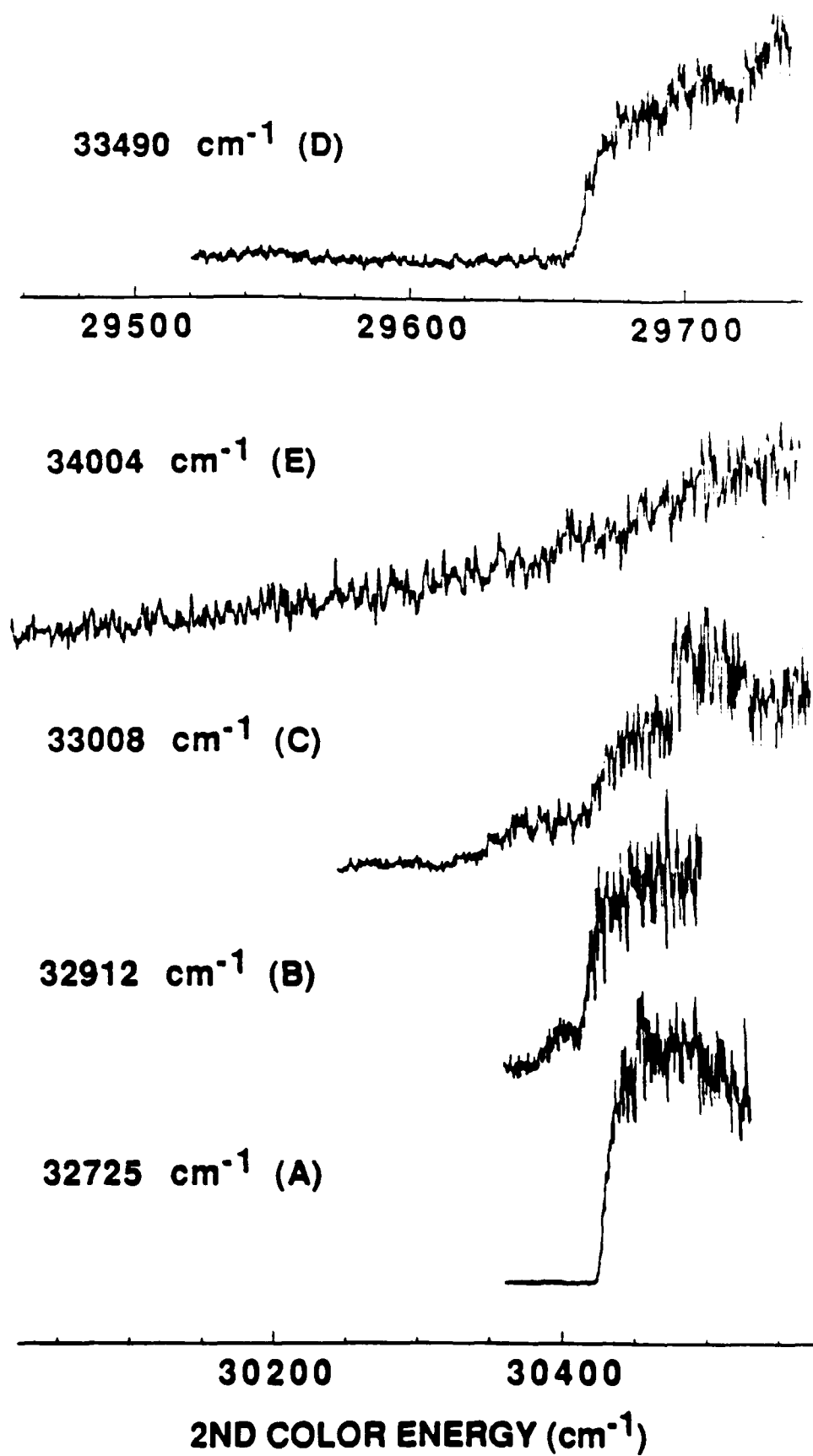


Figure 7

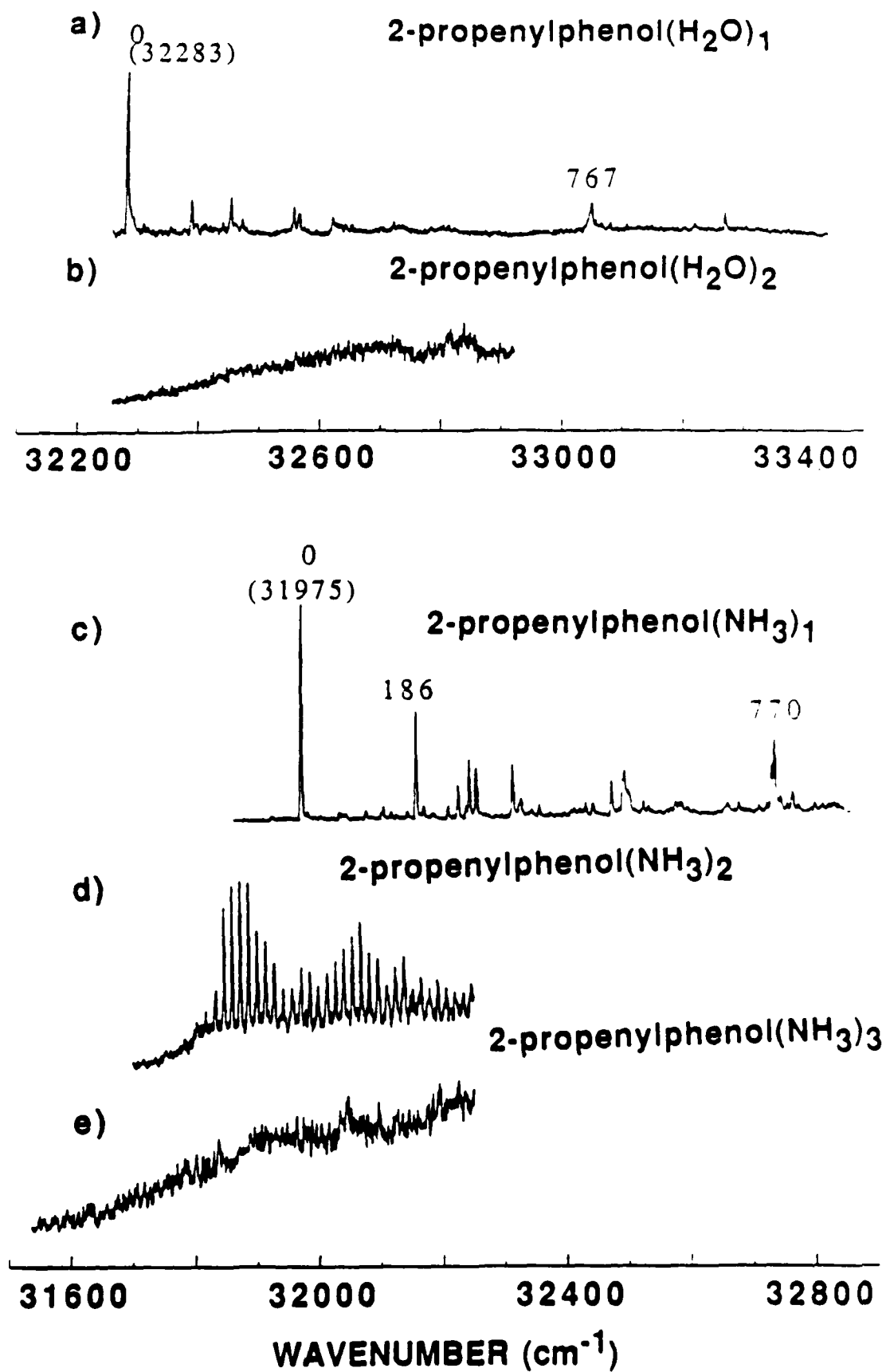
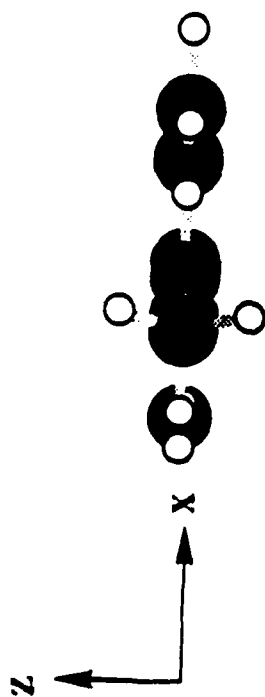
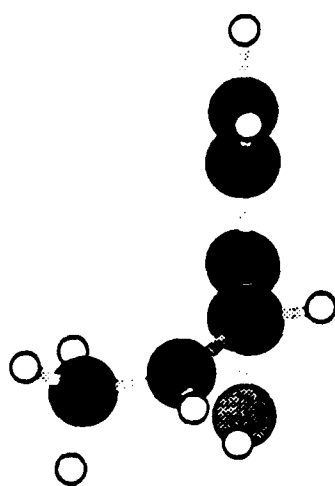


Figure 8

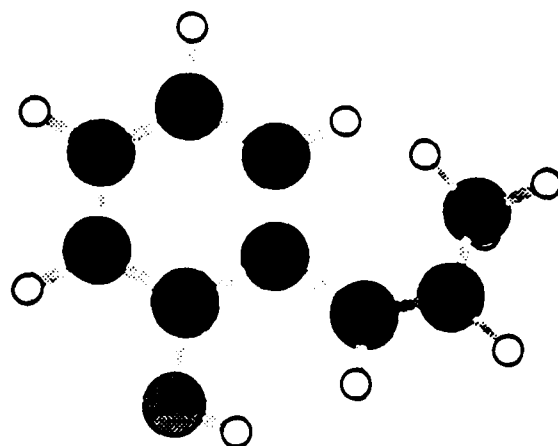
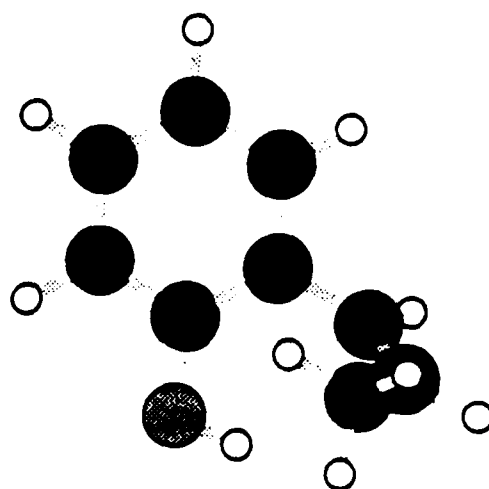
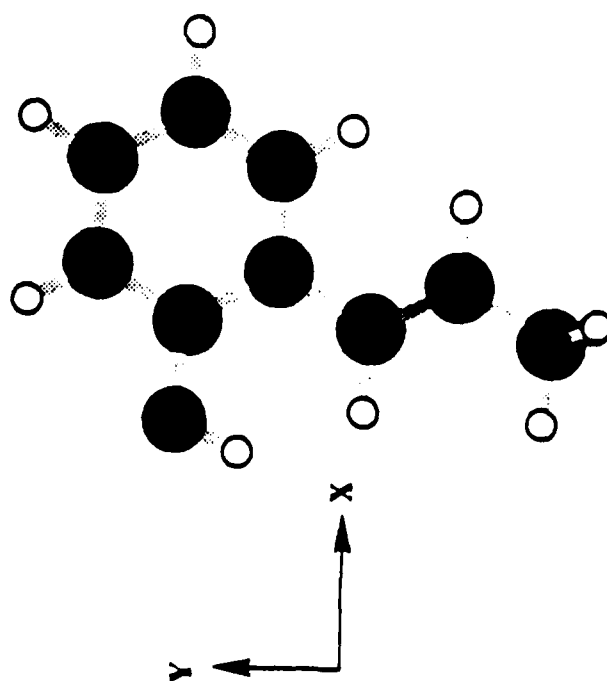
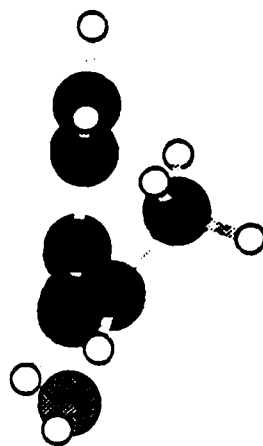
a)



b)



c)



$$\tau_1 = 1^\circ \quad \tau_3 = 180^\circ$$

$$\Delta H_f = -16.3 \text{ kcal/mol}$$

$$\tau_1 = 110^\circ \quad \tau_3 = 177^\circ$$

$$\Delta H_f = -14.0 \text{ kcal/mol}$$

$$\tau_1 = 32^\circ \quad \tau_3 = 173^\circ$$

$$\Delta H_f = -13.8 \text{ kcal/mol}$$

Figure 9



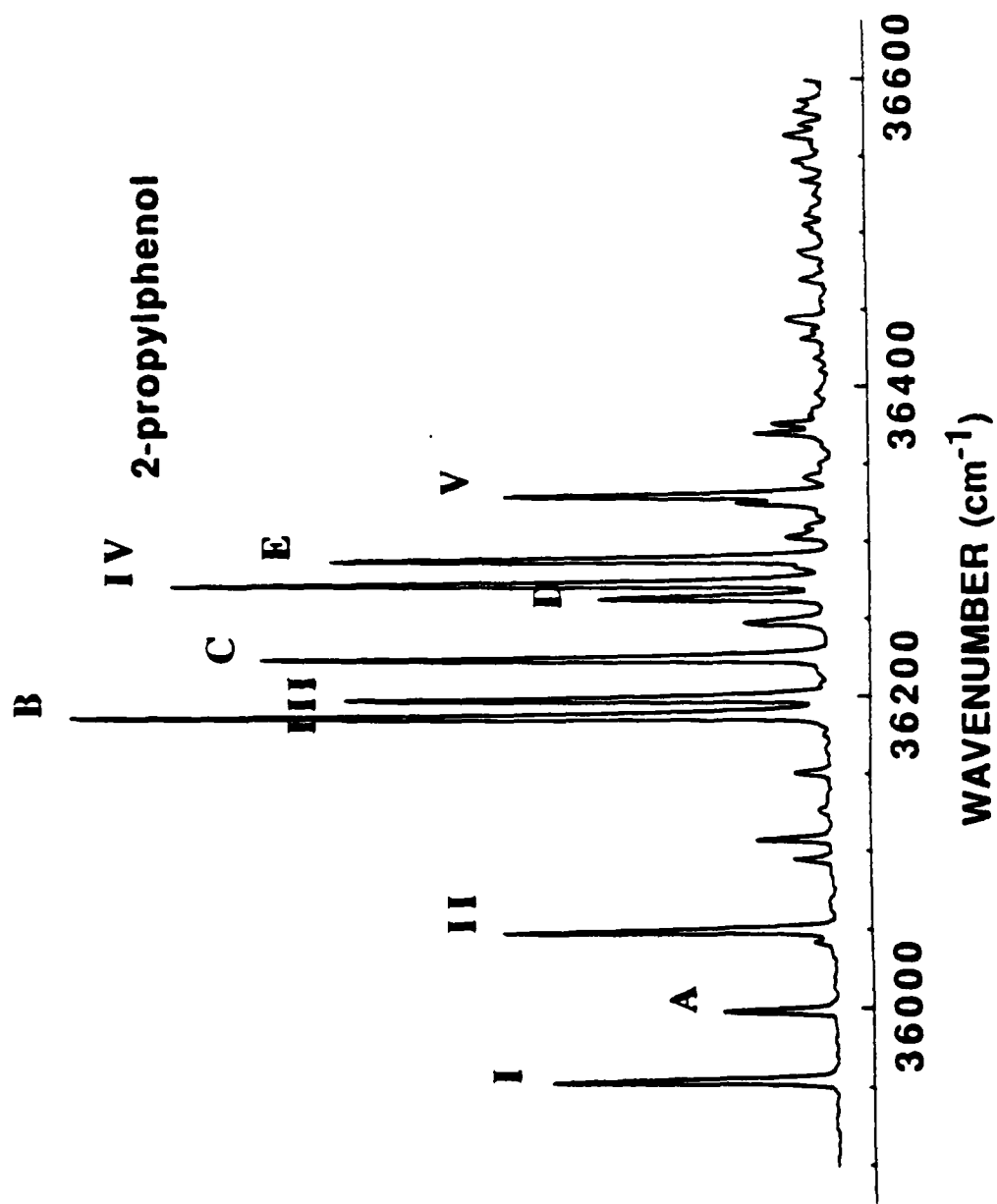


Figure 10

# 2-propylphenol

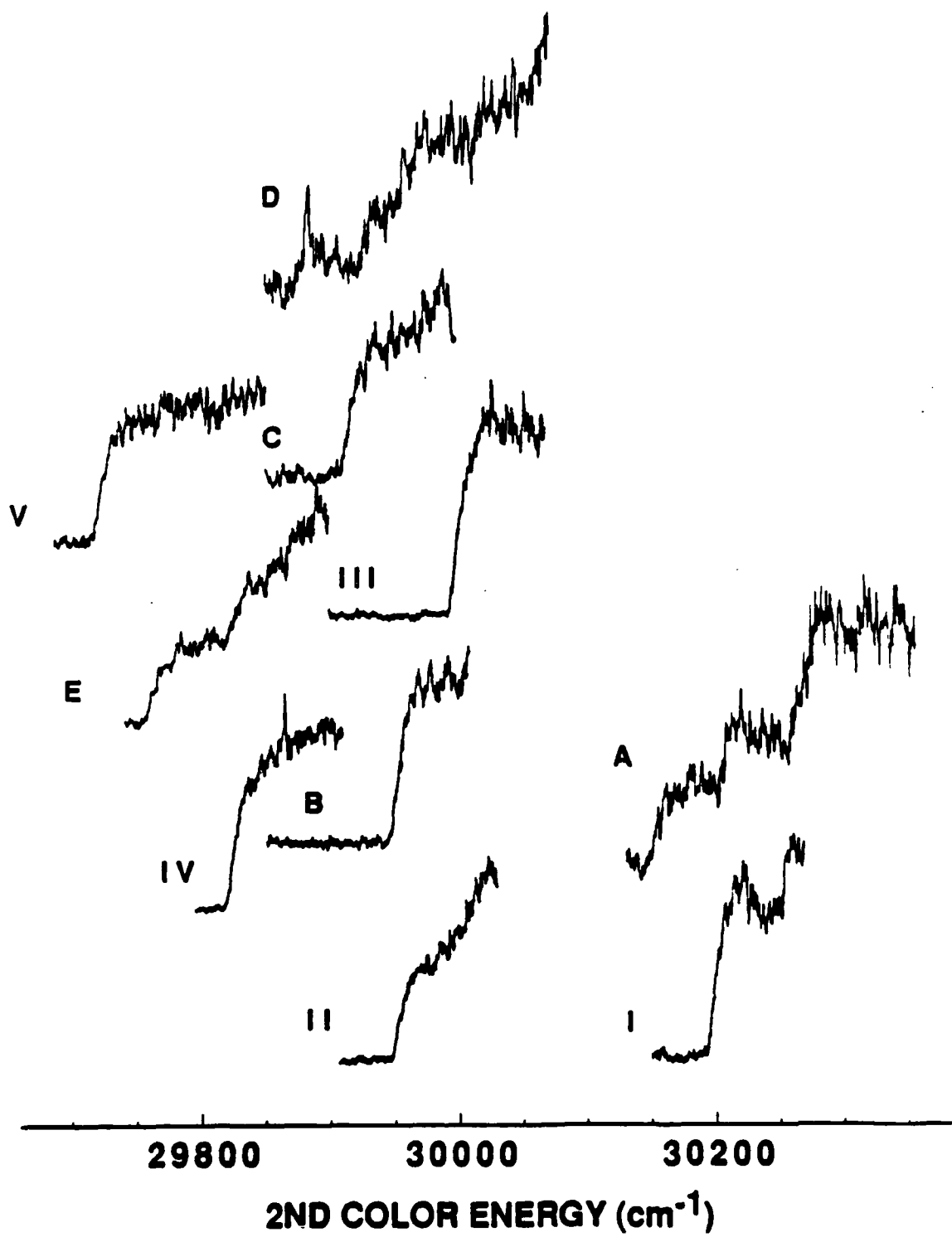


Figure 11

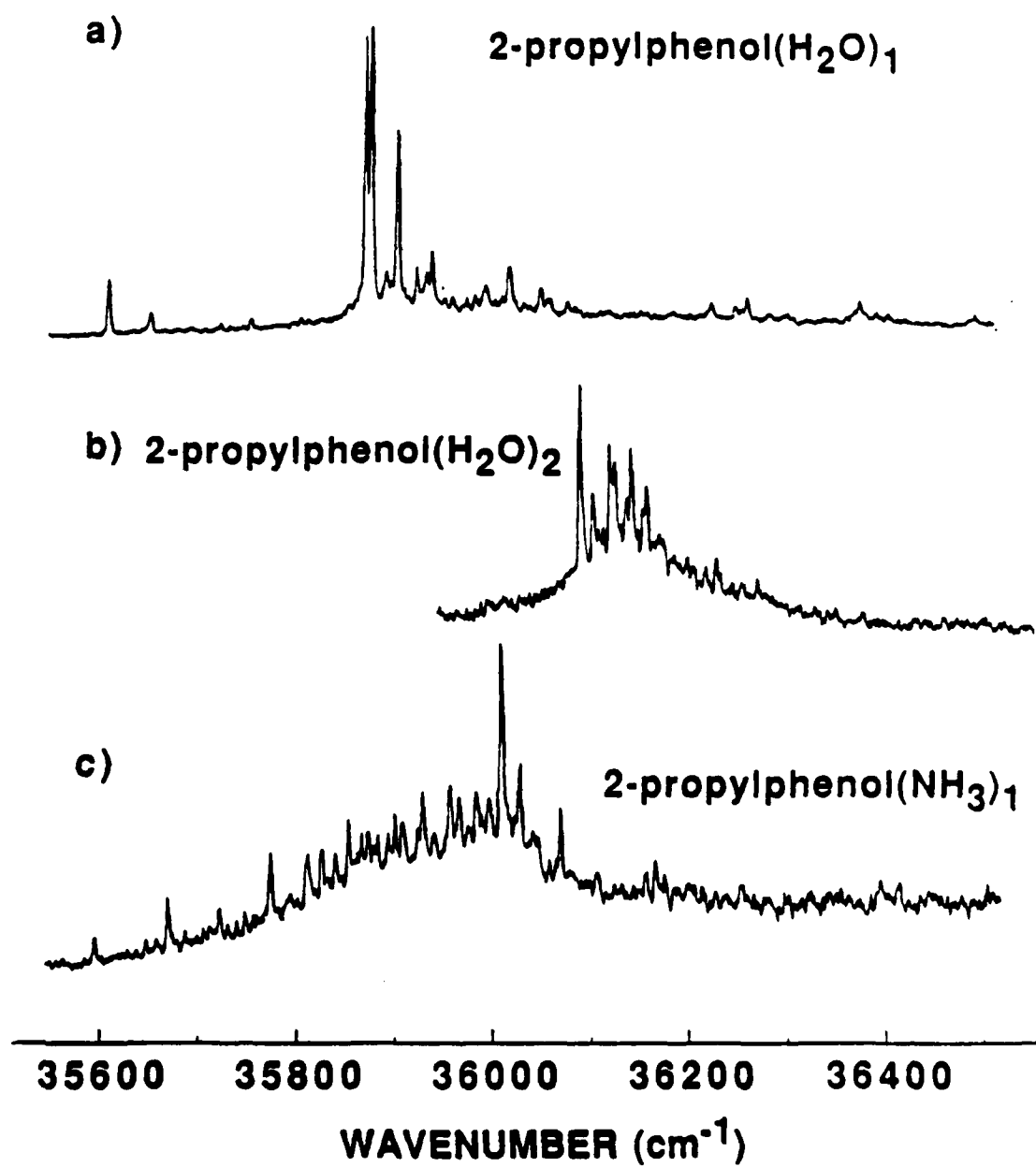


Figure 12